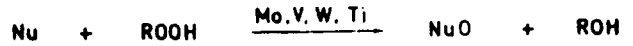
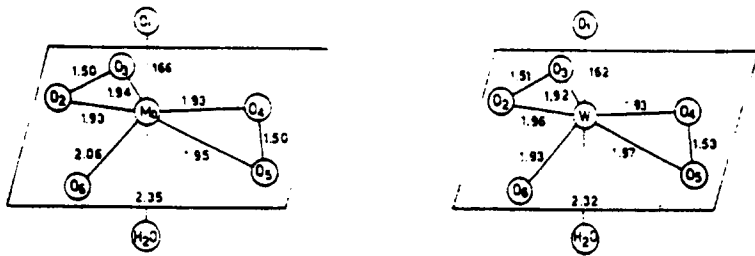
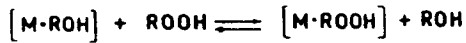
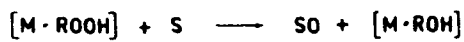
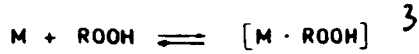


R = H, Alk, Acyl

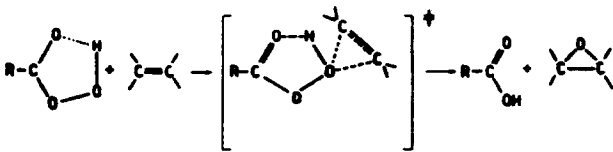
S = R<sub>2</sub>S, R<sub>2</sub>SO, >C=C<, R<sub>3</sub>N, etc



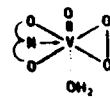
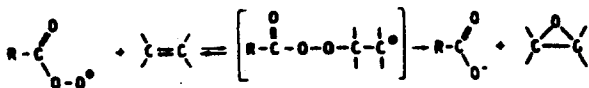
Nu = R<sub>2</sub>S, >C=C<, R<sub>3</sub>N, etc.



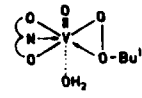
Electrophilic oxygen transfer



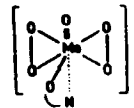
Nucleophilic oxygen transfer



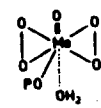
DREW, EINSTEIN, 1973



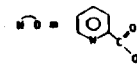
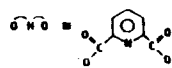
MIMOUN, et al. 1983



5 JACOBSON, TANG, MARES, 1978



LE CARPENTIER, SCHLUPP, WEISS, 1972



PO = [(CH<sub>2</sub>)<sub>3</sub>N]<sub>3</sub>P=O

Oxidant

RCO<sub>3</sub>H    V-TBHP    Mo-TBHP    MoO<sub>3</sub>L



1    ~3 x 10<sup>-4</sup>    ~4 x 10<sup>-3</sup>    ~3 x 10<sup>-2</sup>

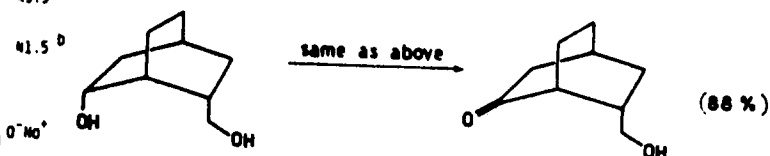
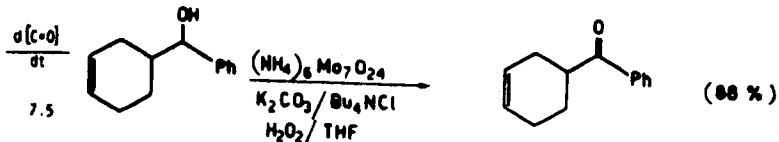
Solvent    benzene    benzene    benzene    DCE

RELATIVE RATES AT ~25°

	-S-	-S- O	Solvent
RCO <sub>3</sub> H	1	0.3	Dioxane
RCO <sub>3</sub> H	1	2 x 10 <sup>-2</sup>	EtOH-Dioxane 70-30
RCO <sub>3</sub> <sup>-</sup>	~0	9 x 10 <sup>-3</sup>	EtOH-Dioxane 70-30
V-TBHP	1	2 x 10 <sup>-2</sup>	EtOH
Mo-TBHP	1	7 x 10 <sup>-4</sup>	EtOH
Mo O <sub>3</sub>	1	1 x 10 <sup>-2</sup>	EtOH

Mo(VI)-CATALYZED OXIDATION OF CYCLOHEXANOL  
(Cyclohexanol as solvent)

#	[BASE] /[Mo(VI)]	$-\frac{d[H_2O_2]}{dt}$
1	0	2.7
2	1	52.8
3	2	50.8
4 (in $CH_3CN$ ) <sup>d</sup>	0	3.7



Mo(VI) =  $MoO_2(acac)_2$ ,  $1 \cdot 10^{-4}$  M ;  $H_2O_2 = 4.5 \cdot 10^{-3}$  M ; BASE =  $cC_6H_{11}O^-Mo^+$

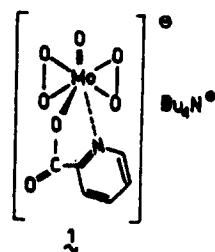
a - Cyclohexanol 1 M in  $CH_3CN$  ; b - Initial rate

OXIDATION OF ALCOHOLS UNDER PHASE-TRANSFER CONDITIONS AT 75°C

Alcohol	Catalyst	t (min)	Product (yield %)
Cyclohexanol	Mo	150	Cyclohexanone (88)
"	W	50	" (-85)
Menthhol	Mo	210	Menthhone (-85)
"	W	120	" (88)
Borneol	Mo	120	Borneone (-85)
"	W	45	" (-85)
2-octanol	W	75	2-octanone (-85)
Benzyl alcohol	W	30	Benzaldehyde (85)

DCE (10 ml)- water (1 ml), pH = 3 (Mo), 1.4 (W)

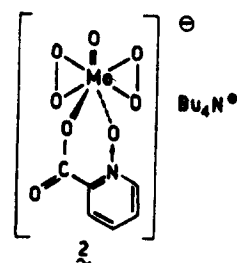
Substrate :  $H_2O_2$  - catalyst : Aliquat 336 = 1 : (2-6) : 0.1 : 0.5



Cyclohexanol  
M  
1.2  
0.4

Oxidant

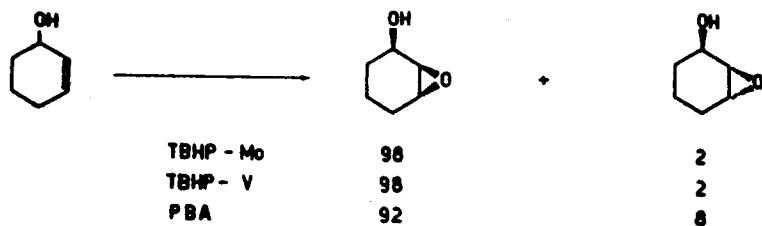
1  
2



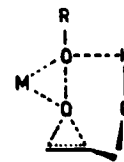
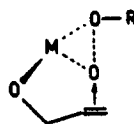
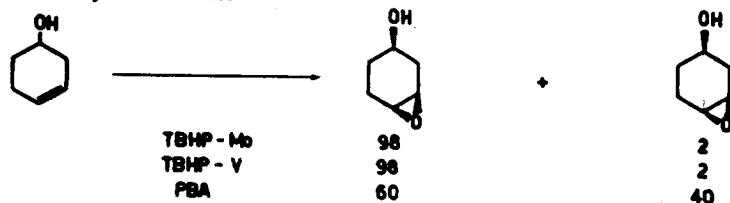
$10^4 k_1 \cdot S^{-1}$   
0.52  
45.0

Solvent  
DCE

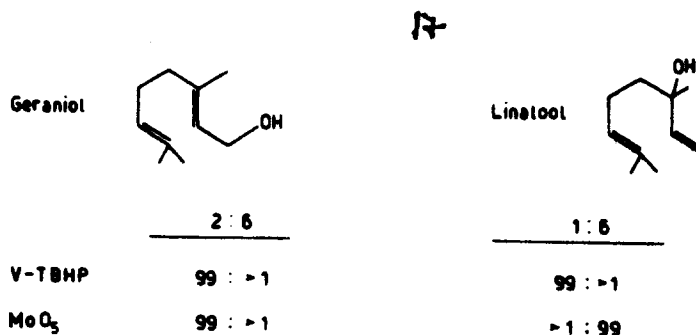
Allylic Alcohols



Homoallylic alcohols



$R_2S$	Me-TBHP		Me-H <sub>2</sub> O <sub>2</sub>	
	$E_A$ (Kcal/mol)	log A	$E_A$ (Kcal/mol)	log A
	15.7	15.4	12.9	12.8
	20.4	14.9	18.8	14.2
	28.1	20.2	17.3	13.3



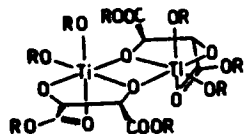
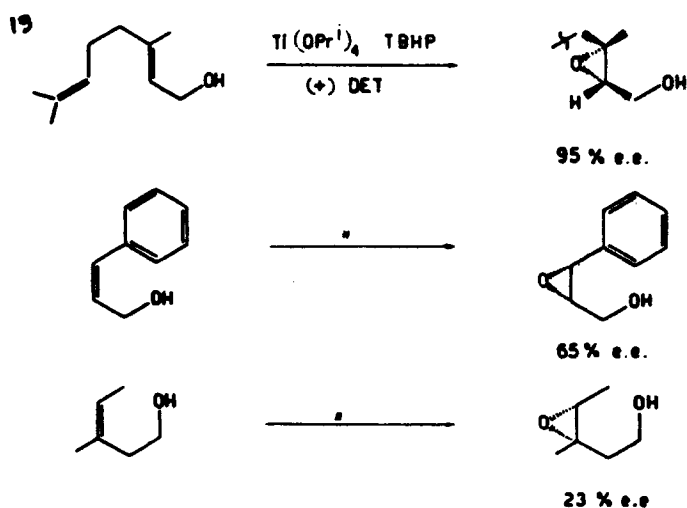
## ENANTIOSELECTIVITY

a) CHIRAL REAGENT (chemical or biochemical)

b) CHIRAL SUBSTRATE

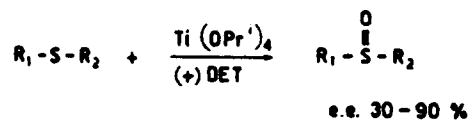
## SUBSTRATE DISCRIMINATION ABILITY

Allylic alcohols ~ Alkenes ~ Sulfides



R = iPr.

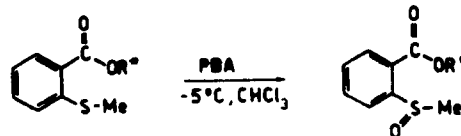
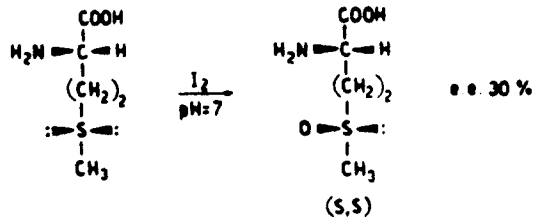
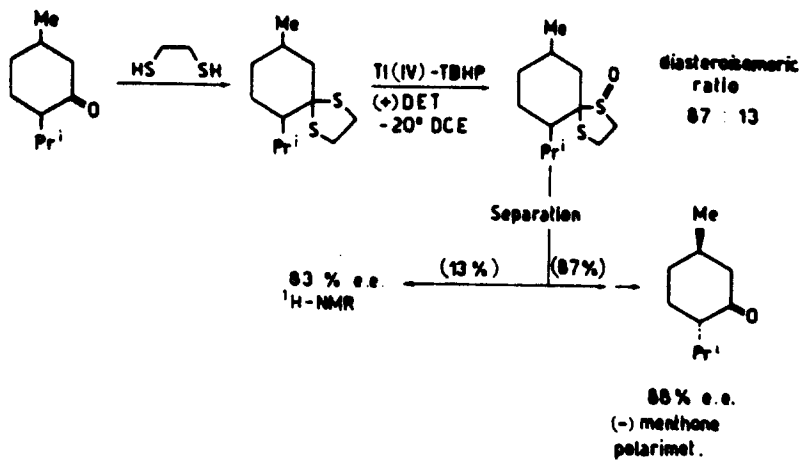
Substrate	Solvent	T °C	e.e. %
	DCE	-20	88
	Toluene	-20	35
	Toluene	-20	14
	DCE	-77	46

e.e. max.  $R_1 = CH_3$ ;  $R_2 = p-CH_3$  Ti (IV) : TBPH : (+) DET : H<sub>2</sub>O

1 : 2 : 4 : ~0

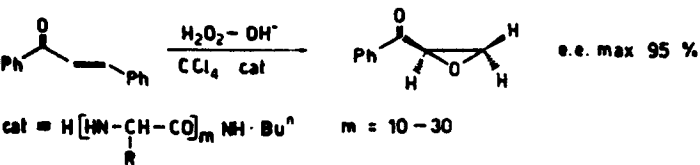
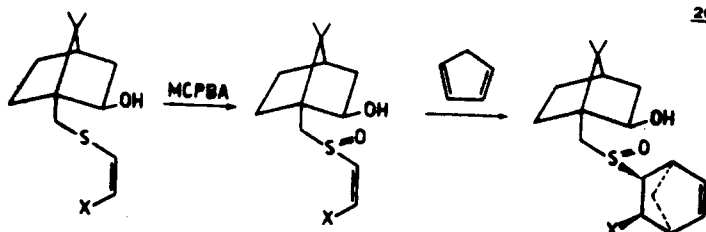
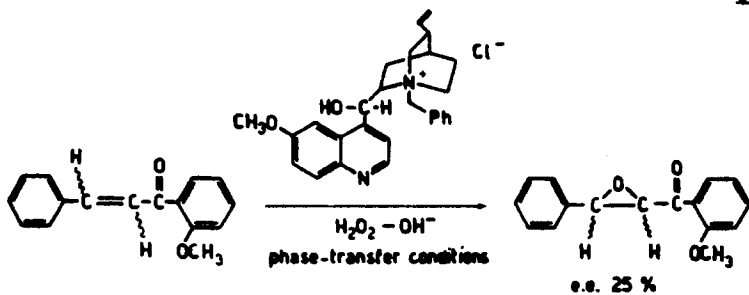
1 : 1.1 : 2 : 1

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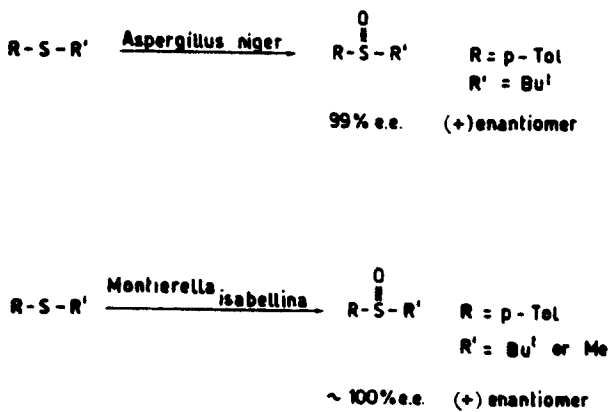
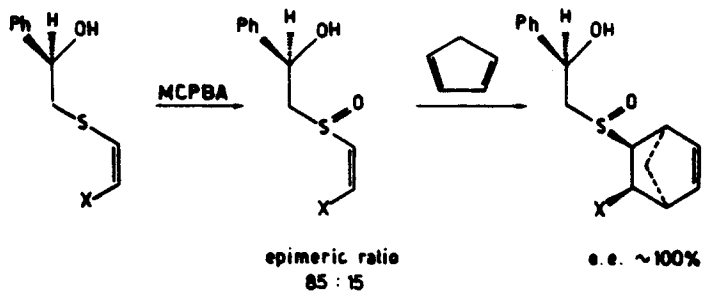
R<sup>+</sup> = (R) Mesityl methyl carbinol

24

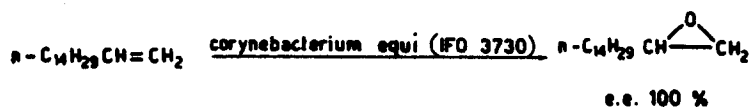
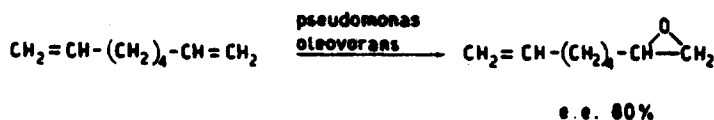


X = SO<sub>2</sub>Ph, SO<sub>2</sub>-C<sub>6</sub>H<sub>4</sub>-Cl, COOMe

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