

Palladium-catalyzed Domino Processes: Serendipity and Rational Design

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Domino Process in Organic Synthesis

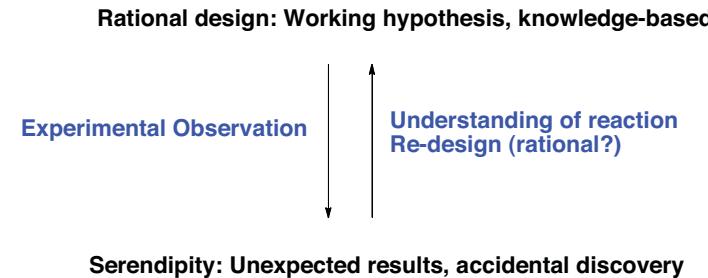
Domino process: a combination of two or more bond-forming reactions under identical conditions *wherein the subsequent reactions result as a consequence of the functionality formed in the previous step.*



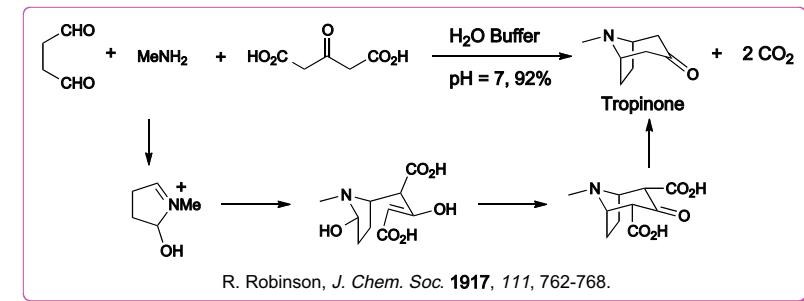
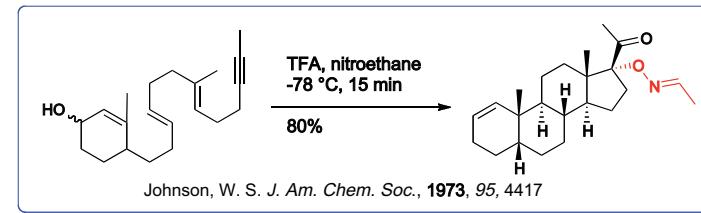
Could be: Uni-molecular (Intramolecular), Bi-molecular and Multi-component

- a) L. F. Tietze, *Chem. Rev.* **1996**, *96*, 115–136;
- b) *Domino Reactions in Organic Synthesis*; L. F. Tietze, G. Brasche, K. Gericke, Eds.; Wiley-VCH, Weinheim, **2006**.
- c) In natural product synthesis: Nicolaou, K. C. *Angew. Chem. Int. Ed.* **2006**, *45*, 7134–7186.

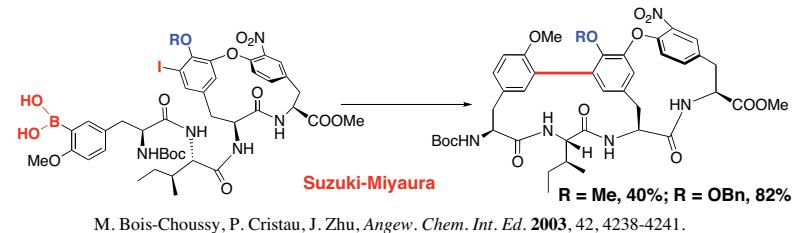
Reaction Discovery: Rational Design and Serendipity



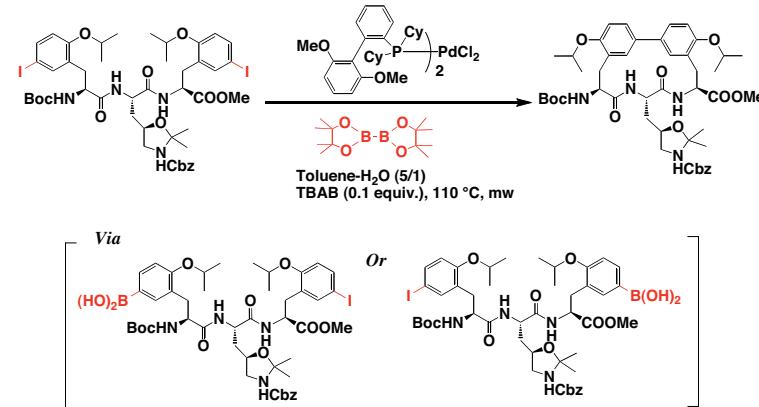
Domino Reactions in Natural Products Syntheses: Classical Examples:



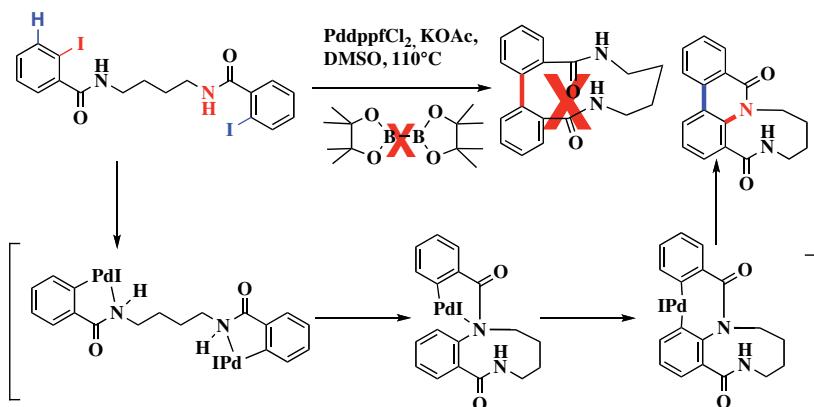
Intramolecular Suzuki-Miyaura Reaction



Total Synthesis of Biphenomycin B

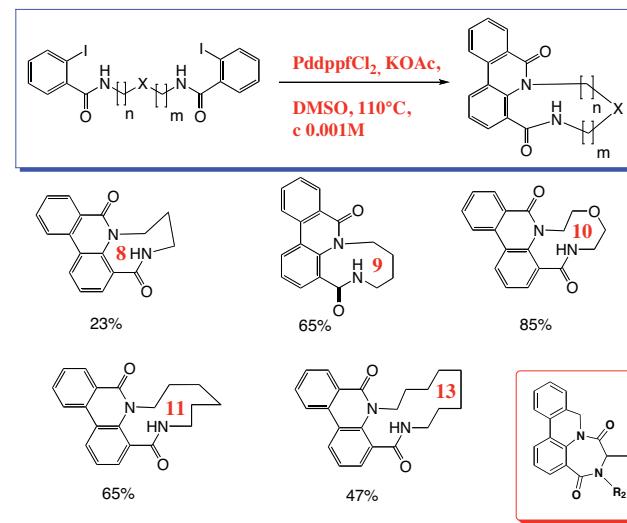


Palladium-Catalyzed Domino Process: Serendipity

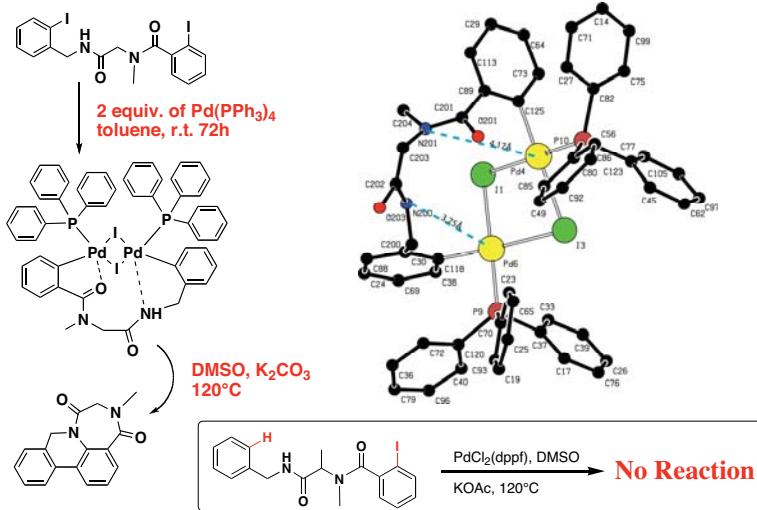


Reviews on Domino process:
L. Tietze, et al. *Chem. Rev.* **1996**, 96, 115-136.
M. Lautens, et al. *Chem. Rev.* **2007**, 107, 174-238.

Azaphenethrene Fused Macrocycles by Domino Intramolecular N-Arylation/C-H Functionalization

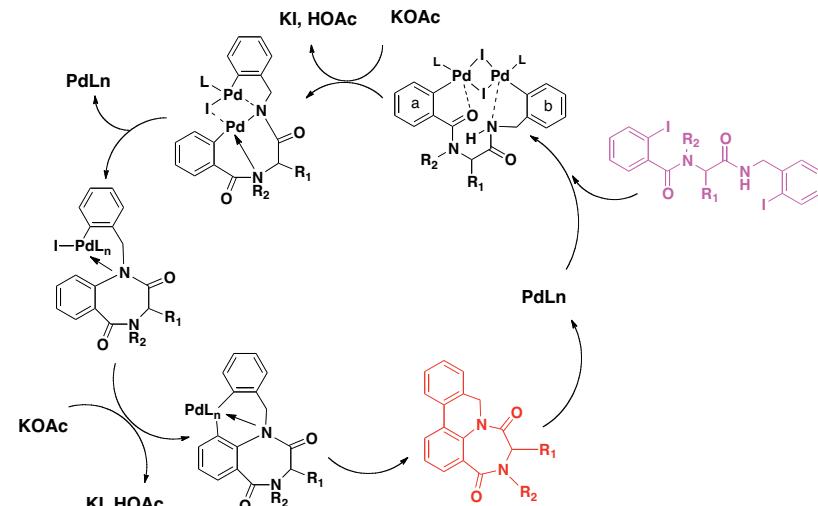


Preparation and Isolation of a Bis-Palladacycle



Salcedo, A.; Neuville, L.; Rondot, C.; Retailleau, P.; Zhu, J. *Org. Lett.* **2008**, *10*, 857-860.
Cuny, G.; Bois-Choussy, M.; Zhu, J. *J. Am. Chem. Soc.* **2004**, *126*, 14474-14484

Proposed Mechanism for the Domino Process



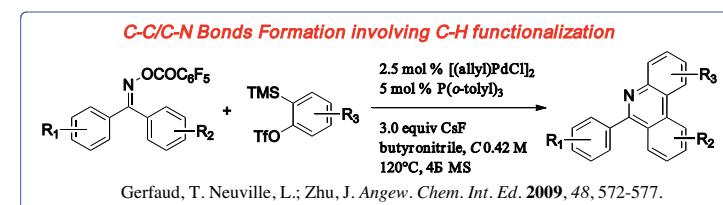
Metal-Catalyzed Domino Process By “DESIGN (?)”

Three key issues needed to be considered:

- **Initiation:**
Carbometallation, Heteronucleomataillation
are ideal
- **Propagation:**
 CO , isonitrile, olefin, allene... good relay
- **Termination:** Any step involving reductive elimination
Cross coupling, Anion capture,
 β -hydride elimination
 $\text{C-H Functionalization}...$

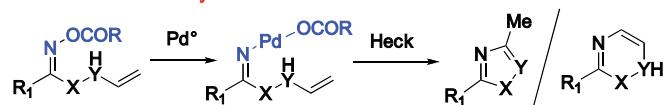
...And Serendipity...

Expanding Narasaka’s O-Acyloxime Chemistry



Chemistry of Acyloxime: Synthesis of Isoquinolines and Phenanthridines

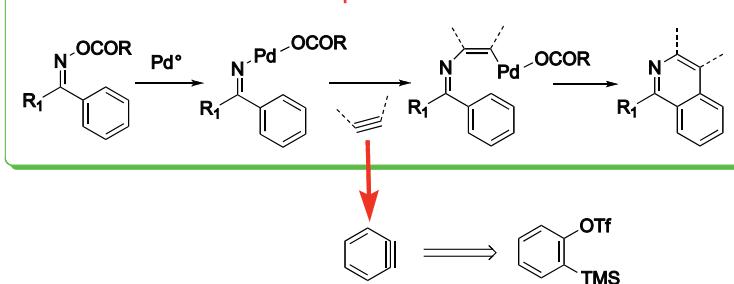
Narasaka's Chemistry



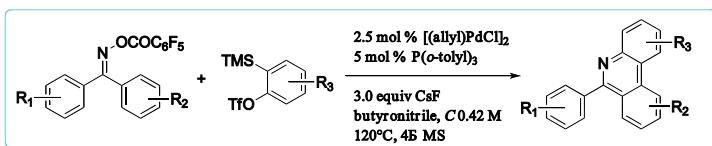
Observation: Intermolecular trapping of =N-Pd-PCOR didn't work

- a) M. Kitamura, K. Narasaka, *Chem. Record* **2002**, 2, 268-277;
- b) K. Narasaka, M. Kitamura, *Eur. J. Org. Chem.* **2005**, 4505–4519.

Domino Intermolecular aminopalladation/C-H functionalisation



Synthesis of phanthridines: Selected Examples



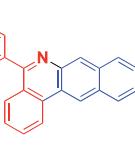
R₁ = R₂ = F, 69%
R₁ = R₂ = CF₃, 43%
R₁ = R₂ = CN, 70%
R₁ = R₂ = Cl, 66%



43%



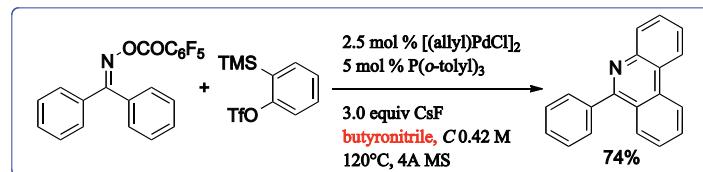
45%



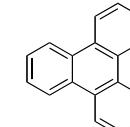
41%

Gerfaud, T. Neuville, L.; Zhu, J. *Angew. Chem. Int. Ed.* **2009**, 48, 572-577.

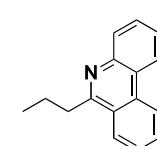
Domino C-C/C-N bonds formation: Condition survey



- **Pd sources:** PdCl₂ leading to trimerization of benzyne
- **Fluoride source:** Tetrabutylammonium triphenyl difluorosilicate (TBAT), instead of CsF, led to the formation of triphenylene
- **Solvent:** C₃H₅CN (bp 115–117 °C) is essential
- **Temperature:** Less efficient at 100 °C; 40% in C₂H₅CN (bp 97 °C)
- **Ligands:** P(o-tolyl)₃ better than dppp and Xphos

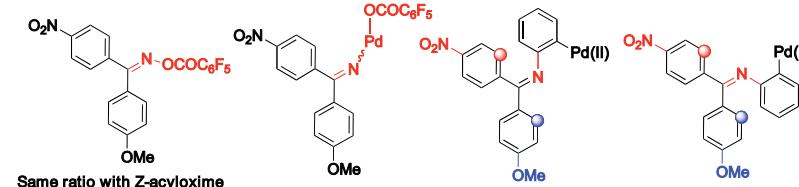
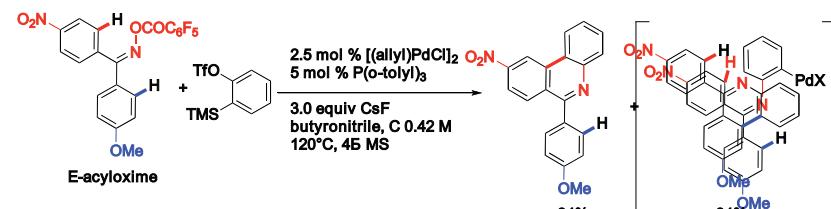


triphenylene

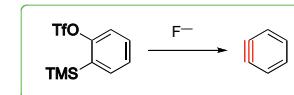


Gerfaud, T. Neuville, L.; Zhu, J. *Angew. Chem. Int. Ed.* **2009**, 48, 572-577.

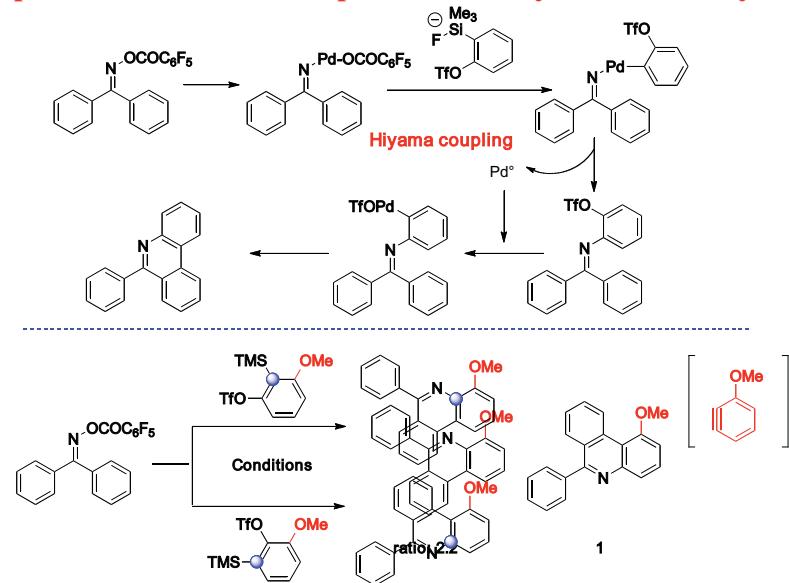
Examples with Mechanistic Implication: Reaction is Insensitive to the Oxime Geometry



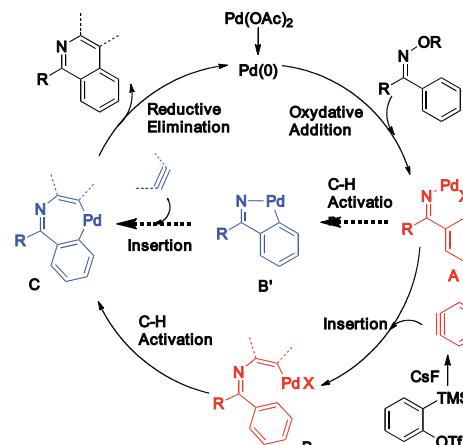
Gerfaud, T. Neuville, L.; Zhu, J. *Angew. Chem. Int. Ed.* **2009**, 48, 572-577.



Examples with mechanistic implication: Benzyne or not benzyne?



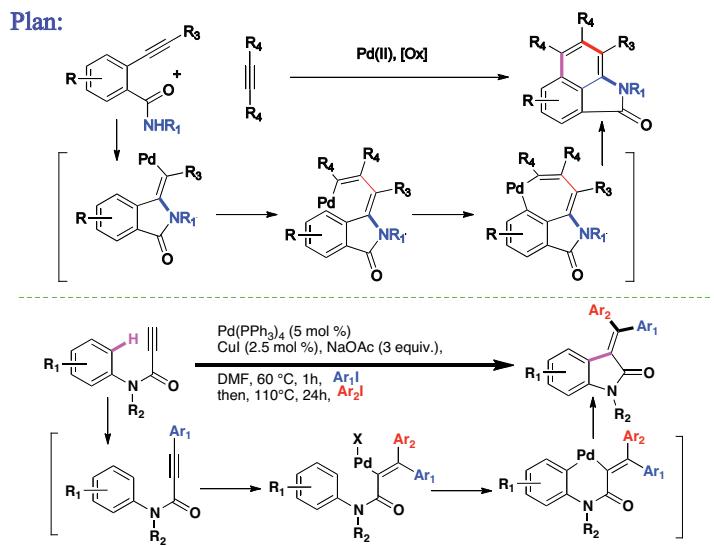
Mechanistic proposal



Key Conclusion: Intermolecular trapping of =N-PdX(II)X by a multiple bond is possible!

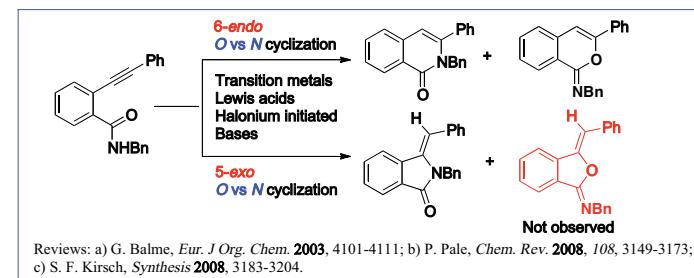
Gerfaud, T. Neuville, L.; Zhu, J. *Angew. Chem. Int. Ed.* **2009**, *48*, 572-577.

Amidopalladation Initiated Domino Process

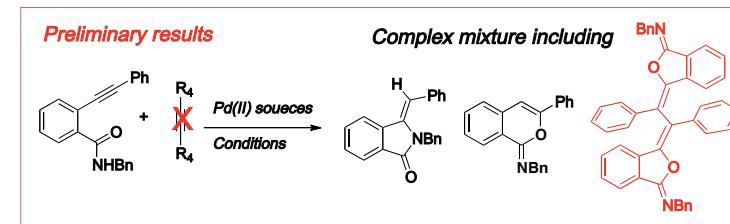


Pinto, A. Neuville, L. Zhu, J. *Angew Chem. Int. Ed.* **2007**, *119*, 3355-3359.

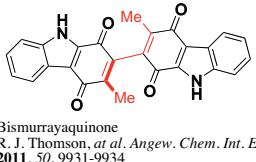
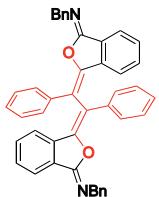
Chemistry of 2-Alkynylbenzamide: Serendipity



Reviews: a) G. Balme, *Eur. J. Org. Chem.* **2003**, 4101-4111; b) P. Pale, *Chem. Rev.* **2008**, *108*, 3149-3173; c) S. F. Kirsch, *Synthesis* **2008**, 3183-3204.



Metal-catalyzed Syntheses of 1,3-Diene



Metal-catalyzed cyclative dimerization processes involving **alkynes**:

Au-catalyzed:

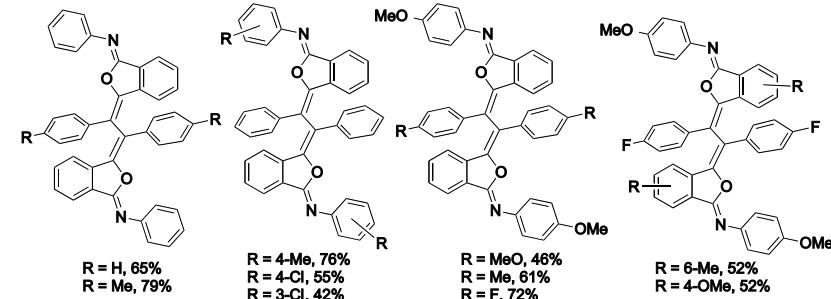
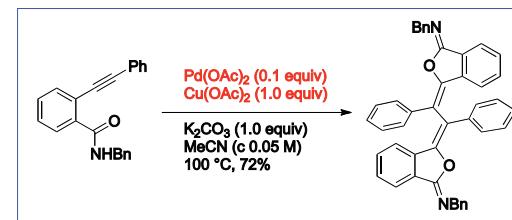
- a) H. A. Wegner, S. Ahles, M. Neuburger, *Chem. Eur. J.* 2008, 14, 11310-11313;
b) K. H. Ahn, *Angew. Chem.* 2011, 123, 11648-11652; *Angew. Chem. Int. Ed.* 2011, 50, 11446-11450;
Pd-catalyzed:
N. Furuichi, H. Hara, T. Osaki, M. Nakano, H. Mori, S. Katsumura, *J. Org. Chem.* 2004, 69, 7949-7959 (mentionned as side products)

Metal-catalyzed cyclative dimerization processes involving **allenenes**,

Pd-catalyzed: S. Ma, *Chem. Eur. J.* 2005, 11, 2351-2356;

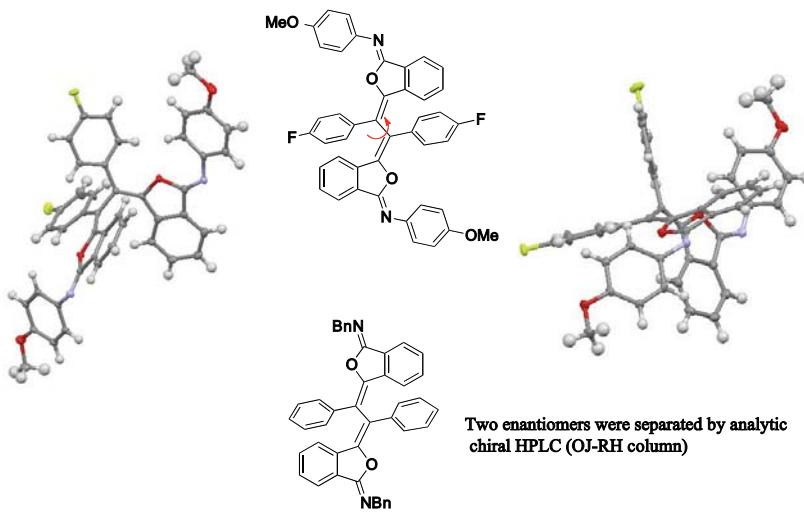
Au-catalyzed: A. S. K. Hashimi, *Eur. J. Org. Chem.* 2006, 1387-1389.

Cyclizative Dimerization: Conditions Survey



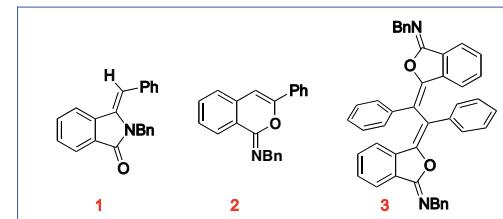
Yao, B.; Jaccod, C.; Wang, Q.; Zhu, J. *Chem. Eur. J.* 2012, 18, 5864-5868.

Axial chirality of the dimer



Synergistic Effect of Pd and Cu

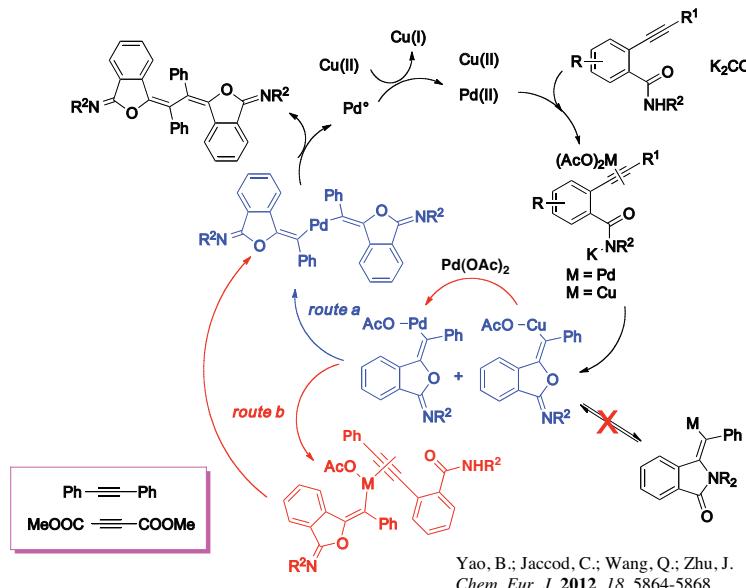
Entry	Pd(OAc) ₂	Cu(OAc) ₂	K ₂ CO ₃	BrNield's
1	-	-	1.0 eq	92% of 4
2	-	-	1.0 eq	Q10% of 3 + 53% of 1
3	-	Pd(OAc) ₂ (0.1 equiv)	1.0 eq	</>10% of 3 - 63% of 1
4	0.1 eq	Cu(OAc) ₂ (1.0 equiv)	1.0 eq	12% + 43% of 2
5	0.1 eq	-	1.0 eq	1% + 35% of 2
6	0.1 eq	Pd(OAc) ₂ (0.1 equiv)	1.0 eq	48% of 3
7	0.2 eq	Pd(OAc) ₂ (0.1 equiv)	1.0 eq	72% of 3
8	0.1 eq	Pd(OAc) ₂ (0.1 equiv)	1.0 eq	100% of 3 + 6% of 1



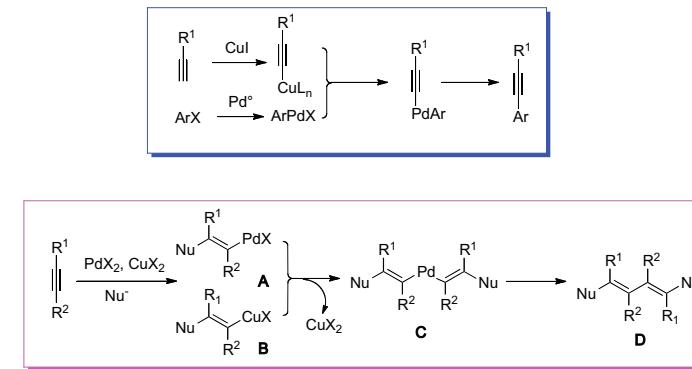
Conclusion:

- Cyclizative dimerization took place only in the presence of both Pd and Cu catalysts.
- Cu(OAc)₂ is not merely acting as an oxidant to convert Pd⁰ back to Pd^{II}.
- Higher concentration produced lower yield of cyclic dimer (results not shown).

Possible Reaction Pathway: Cooperative Effects of Pd and Cu???

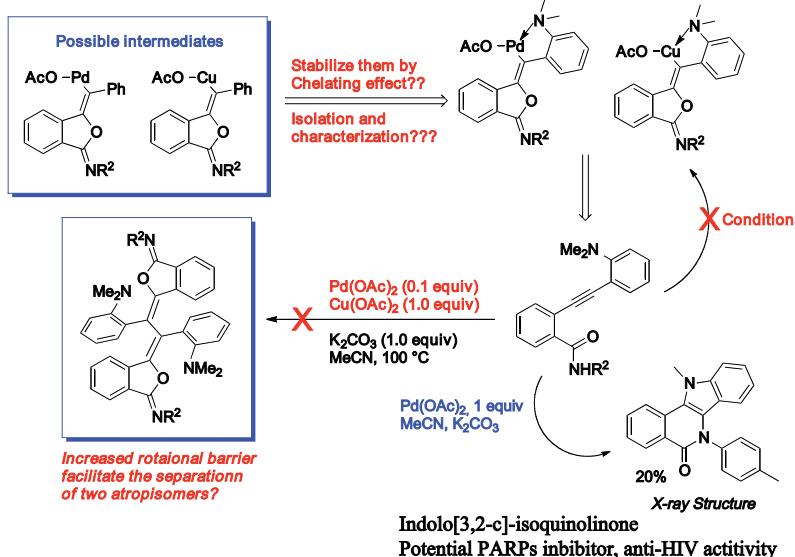


Synergistic Effects of Pd and Cu

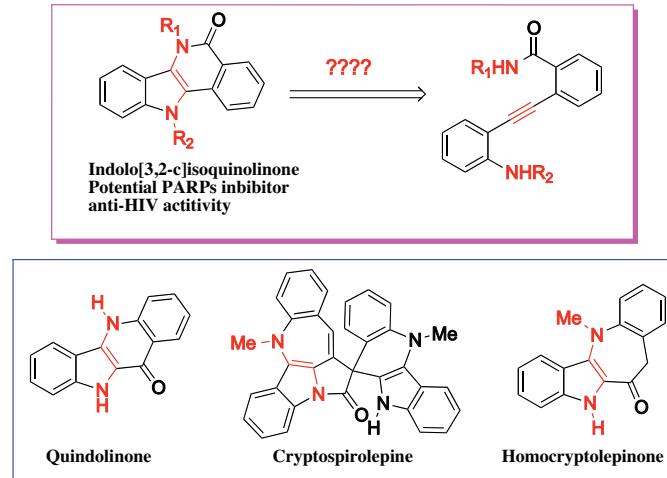


Pd-catalyzed cross coupling of vinylcuprate:
N. Jabri, A. Alexakis, J. F. Normant, *Tetrahedron* **1986**, *42*, 1369-1380.
CuI-accelerated Stille coupling,
L. S. Liebeskind, R. W. Fengl, *J. Org. Chem.* **1990**, *55*, 5359-5364.

Serendipity from a “Rational” Experimental Design

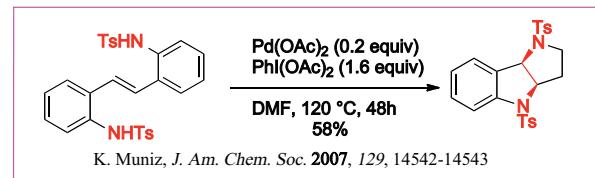
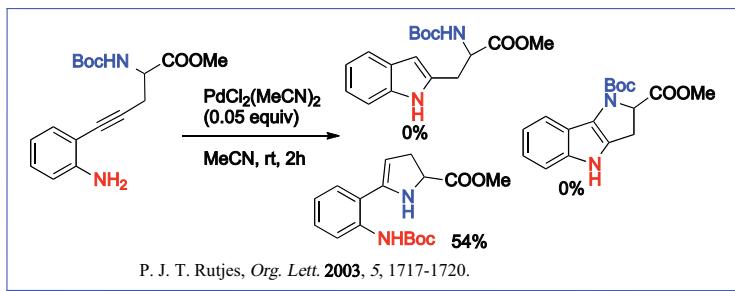


Explore the Serendipity: Diamination of Alkynes?



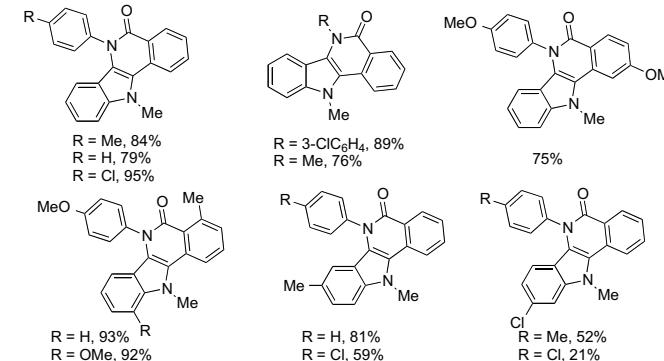
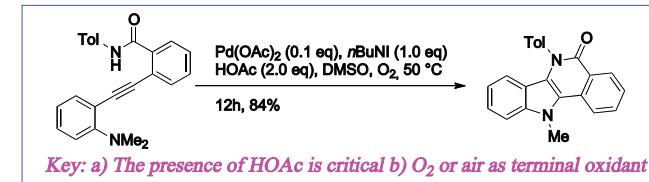
Indolo[3,2-c]isoquinolinone: C. Szabó, *Org. Lett.* **2005**, *7*, 1753 – 1756; L. Li, W. K. S. Chua, *Tetrahedron Lett.* **2011**, *52*, 1574 – 1577. Review : P. Jagtap, C. Szabó, *Nature Rev. Drug Discov.* **2005**, *4*, 421-440.

Diamination of Alkynes: Poorly Studied



For a recent review on nucleopalladation, see: S. S. Stahl, *Chem. Rev.* **2011**, 111, 2981-3019.
For a review on the transition-metal catalyzed diamination, see:
Cardona, F.; Goti A. *Nature Chemistry* **2009**, 1, 269.

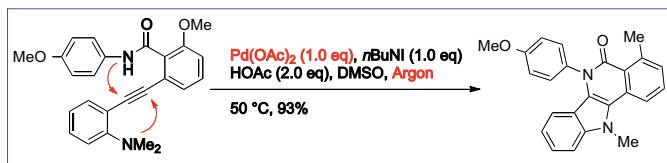
Diamination of Alkynes: Conditions and Scope



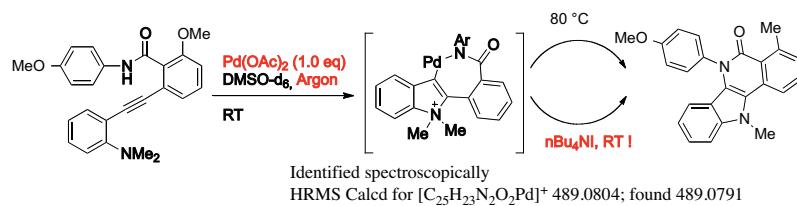
Yao, B.; Wang, Q.; Zhu, J. *Angew Chem. Int. Ed.* **2012**, 51, 5170-5174.

Diamination of Alkynes: Mechanistic Consideration

Question 1: Pd⁰/Pd^{II} or Pd^{II}/Pd^{IV} catalytic cycle



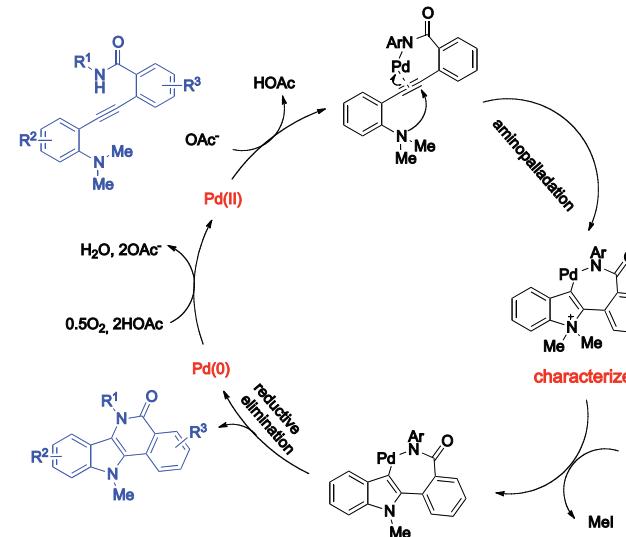
Question 2: Order of cyclization sequence and role of iodide



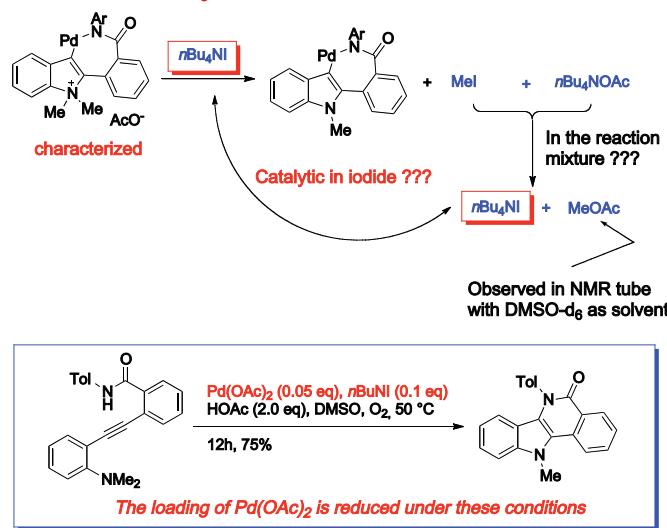
Conclusion:

- a) reaction went through a Pd⁰/Pd^{II} or Pd^{II}/Pd^{IV} catalytic cycle
- b) Indole formation preceeded the quinolinone formation
- c) AcO⁻ can also effect the N-demethylation, but Iodide is much more efficient.

Diamination of Alkynes: A Possible Mechanism

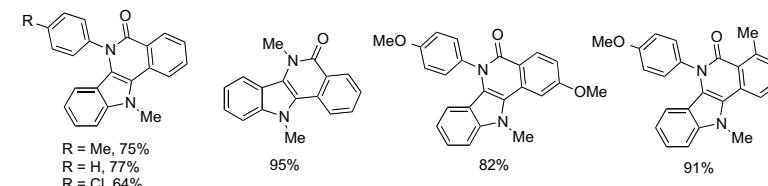
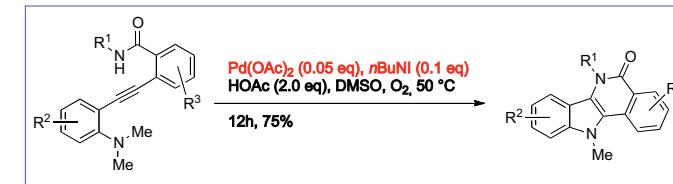


Diamination of Alkynes: Further Mechanistic Consideration

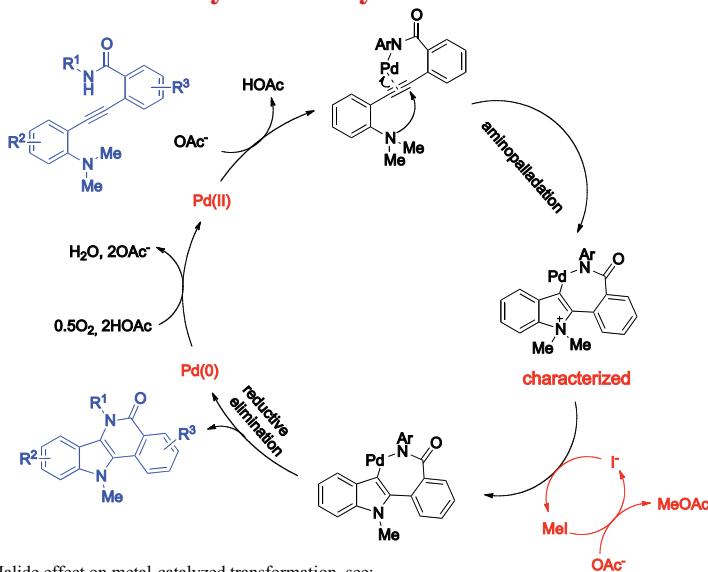


Yao, B.; Wang, Q.; Zhu, J. *J. Angew Chem. Int. Ed.* **2012**, *51*, 5170-5174.

Diamination of Alkynes with Catalytic Amount of Iodide: Examples



Diamination of Alkynes: Catalytic Turnover of Iodide Anion



Halide effect on metal-catalyzed transformation, see:
K. Fagnou, M. Lautens, *Angew. Chem. Int. Ed.* **2002**, *41*, 26-47.

Reaction Discovery: Rational Design and Serendipity

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Dr. Guylaine Cuny, Dr. Artur Pinto

EPFL, Switzerland:

Dr. Qian Wang
Dr. Bo Yao, T. Piou

