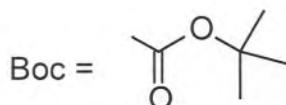
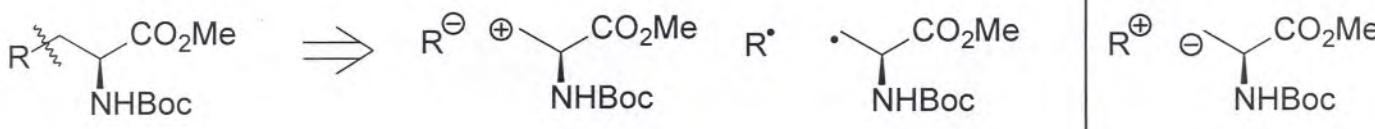
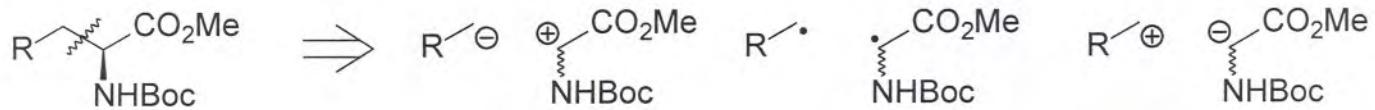


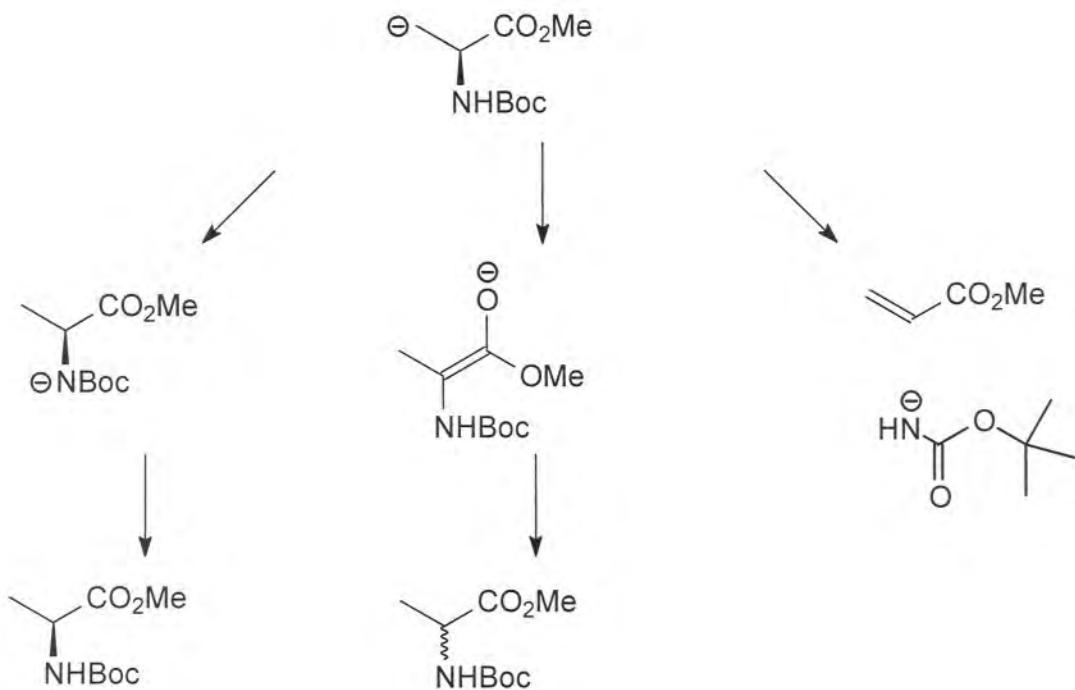
Understanding the Stability and Reactivity of Functionalised Organozinc Reagents: Applications to Amino Acid and Cyclic Peptide Synthesis

R F W Jackson
University of Sheffield

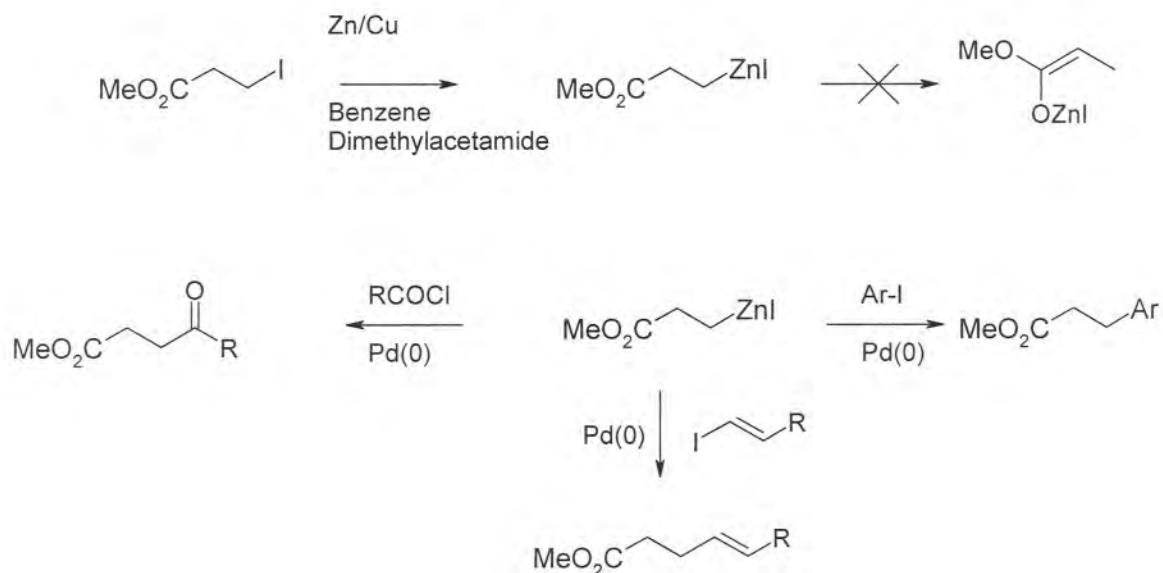
Amino Acid Disconnections



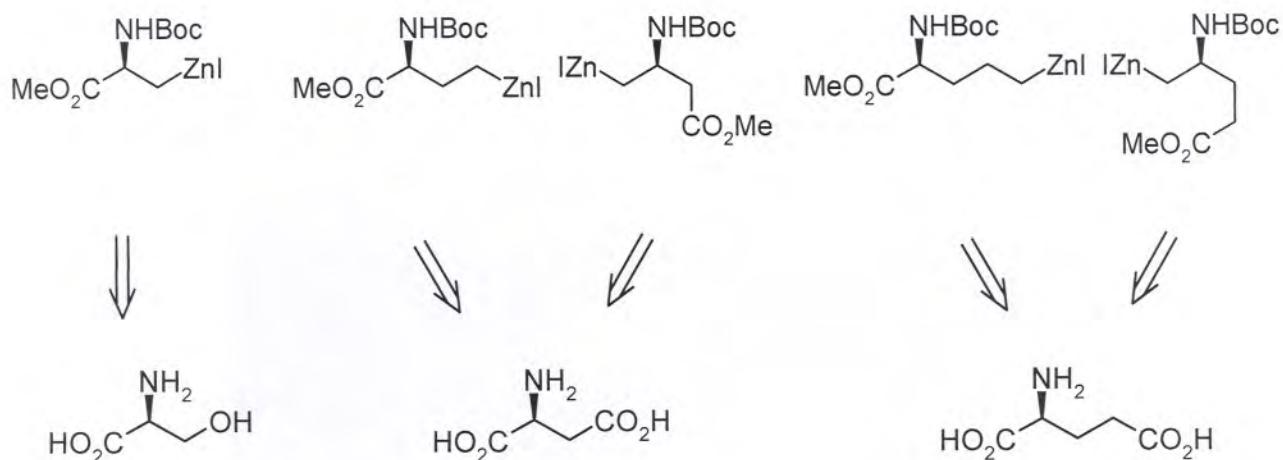
Instability of β -Amino Anions



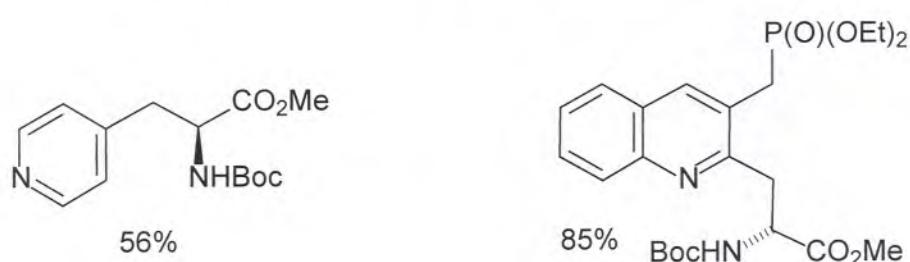
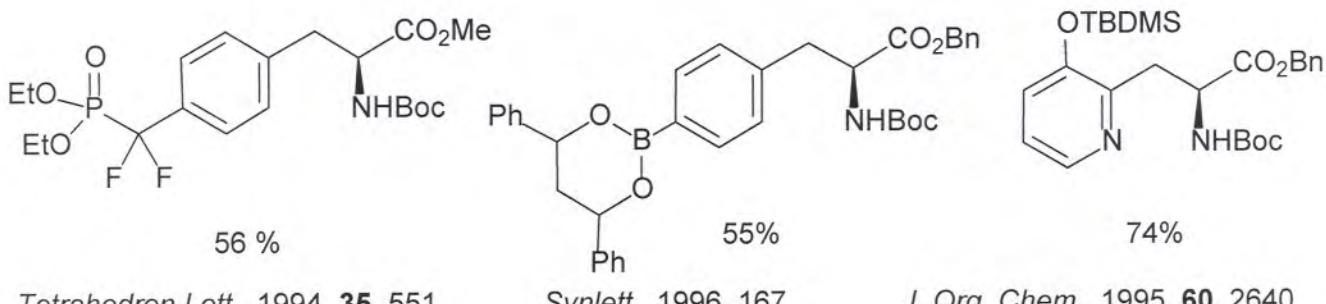
Reactions of Simple Organozinc Reagents



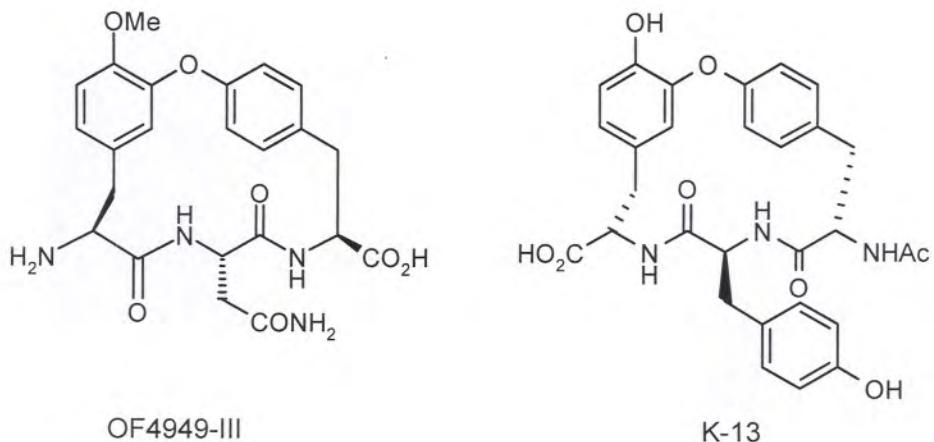
Amino Acid Derived Zinc Reagents



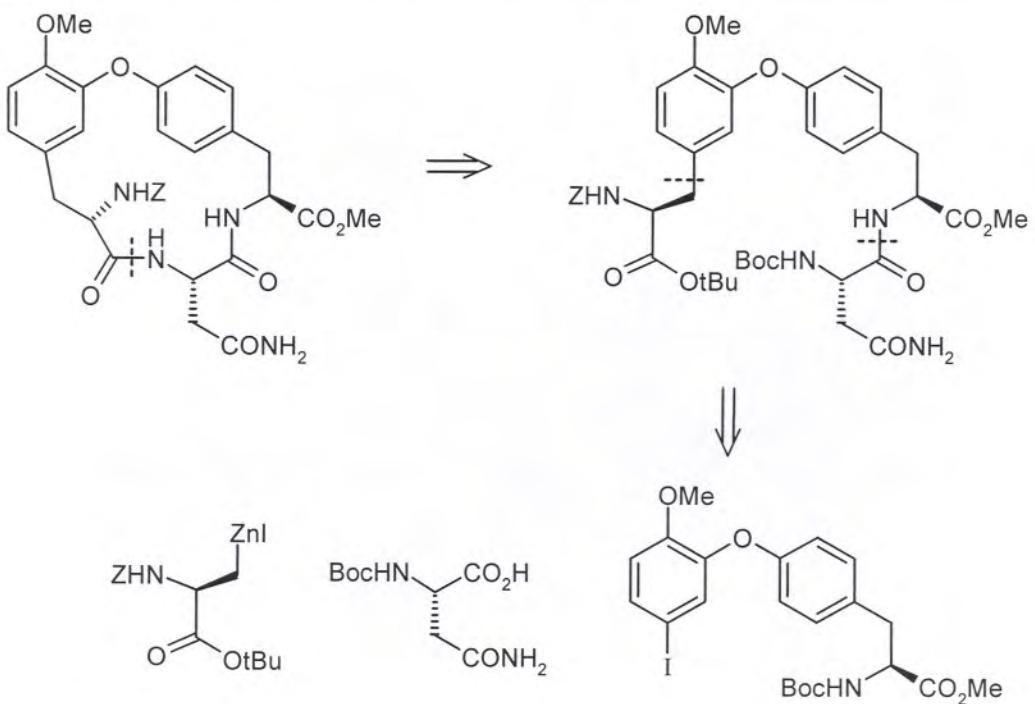
Synthesis of Phenylalanine Derivatives



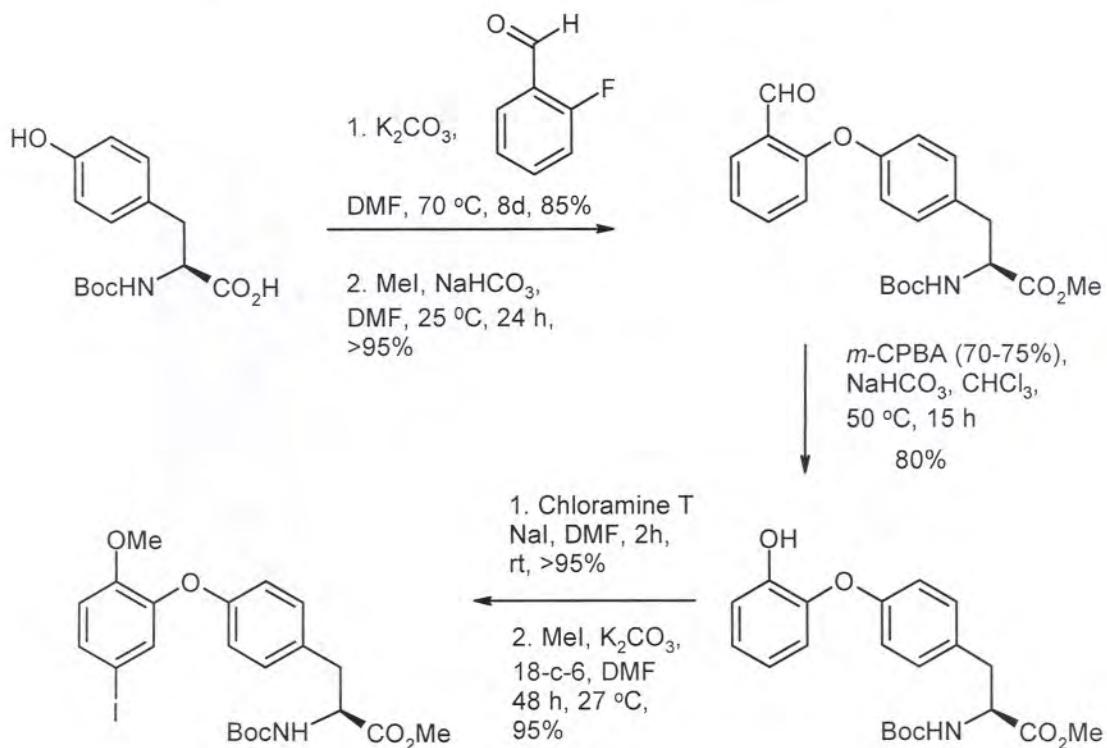
Synthesis of Macrocyclic Tripeptides



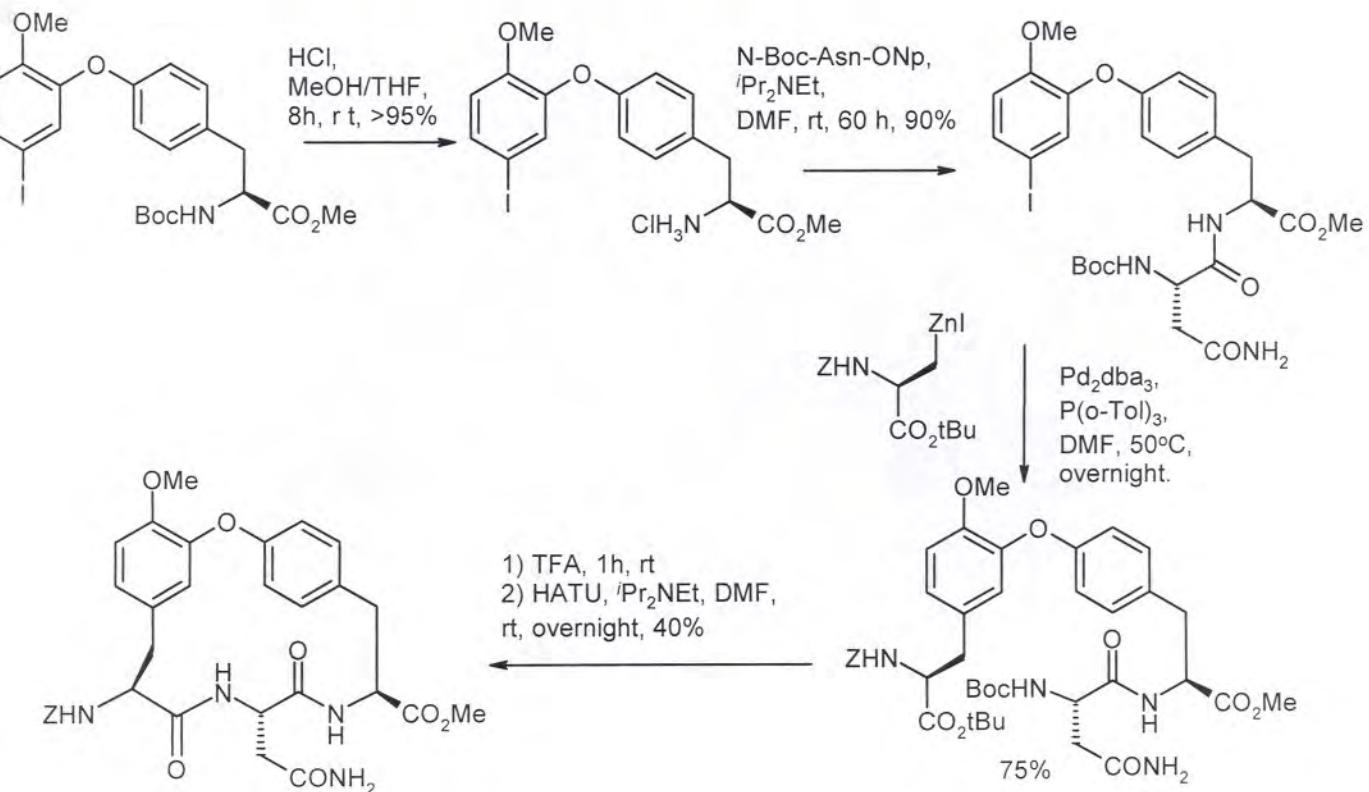
Retrosynthesis for OF4949-III



Synthesis of *O*-Aryltyrosine

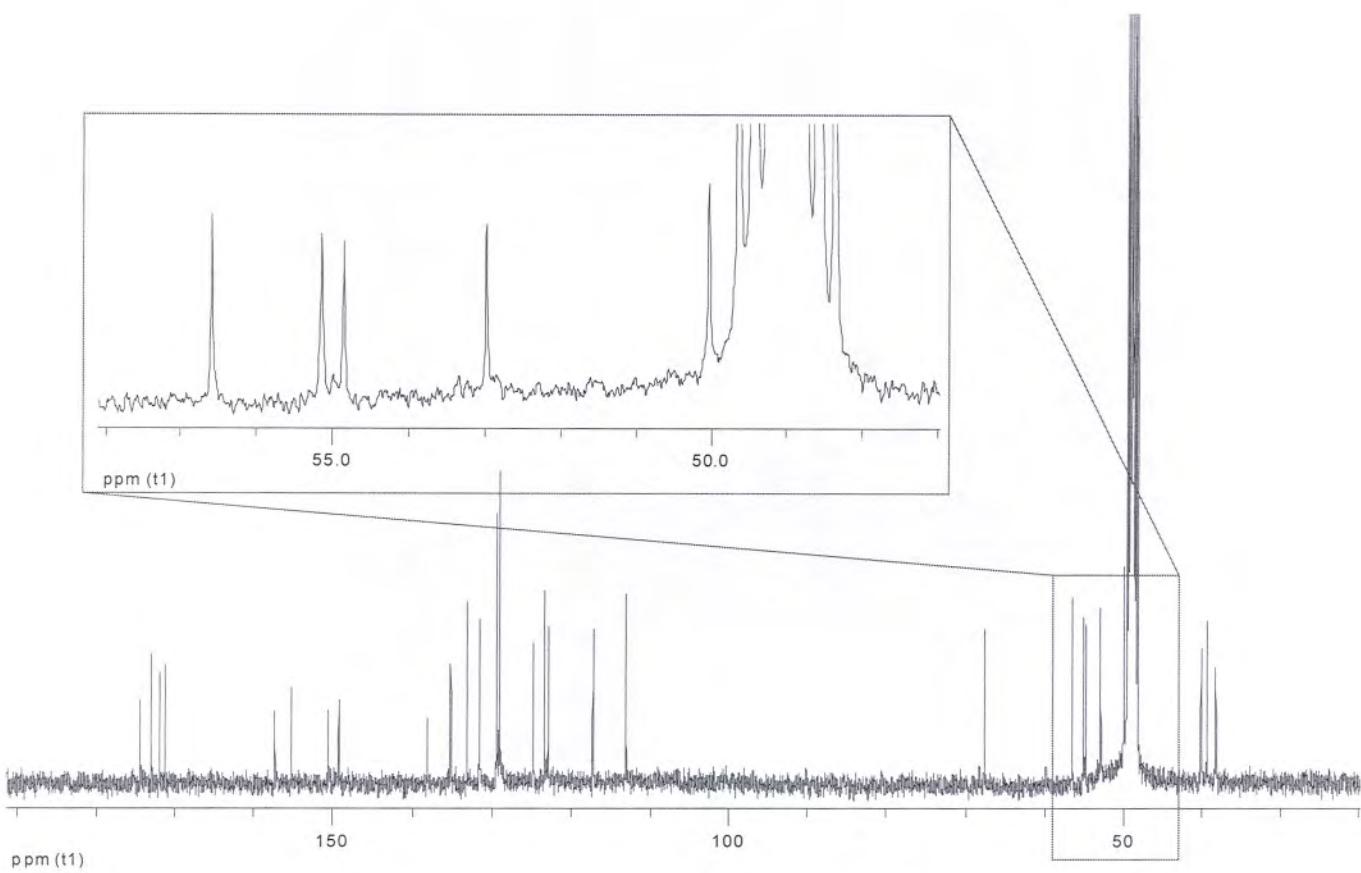
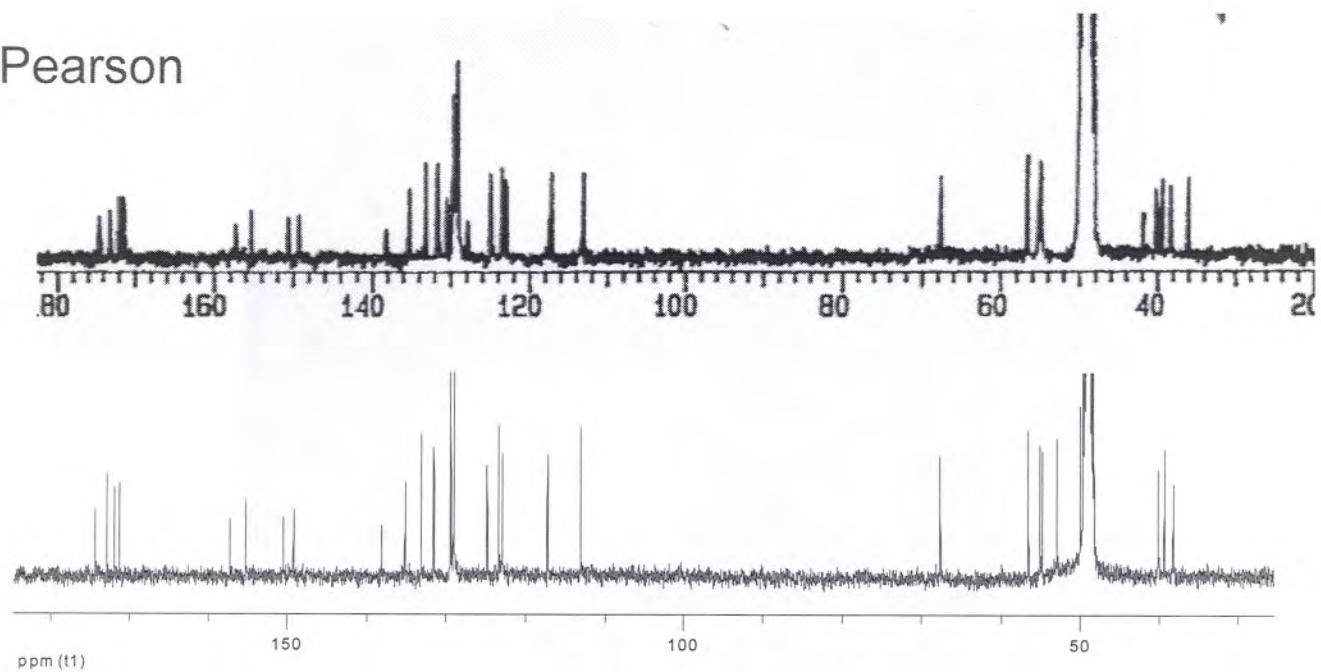


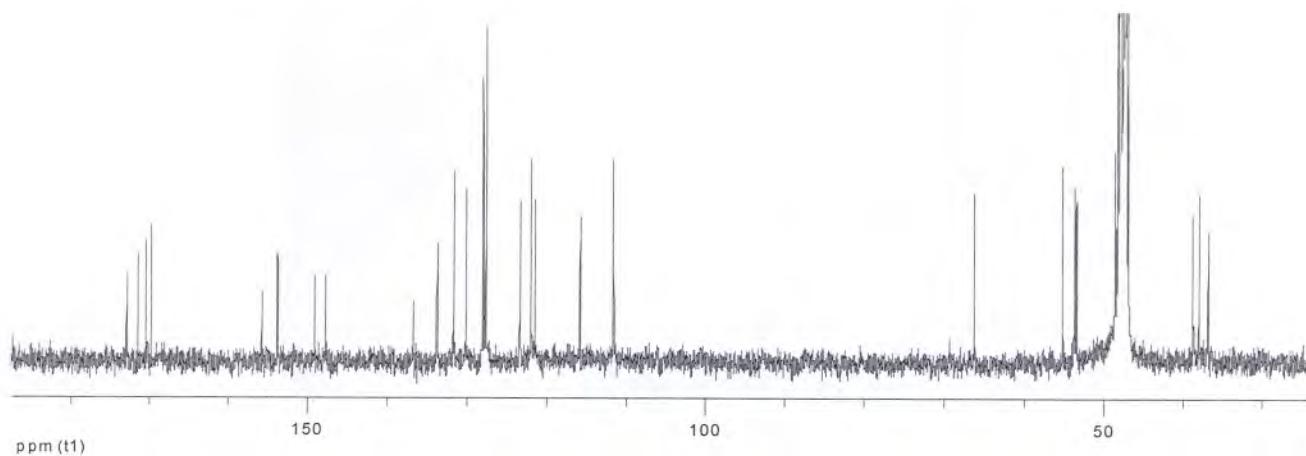
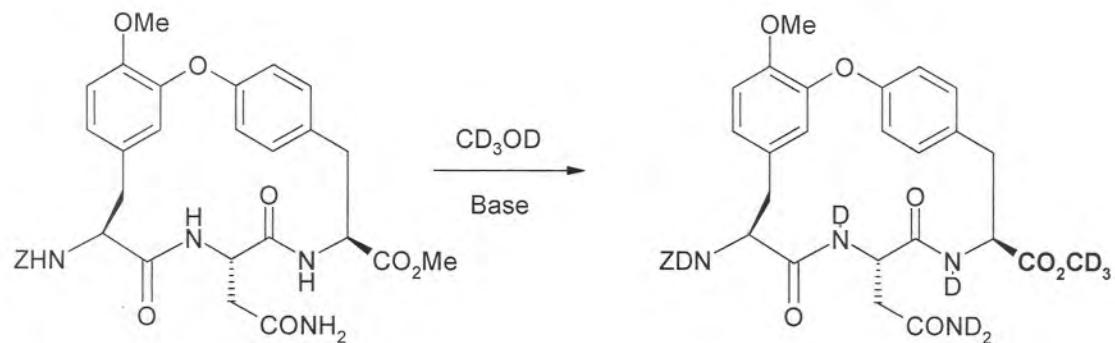
Synthesis of OF4949-III



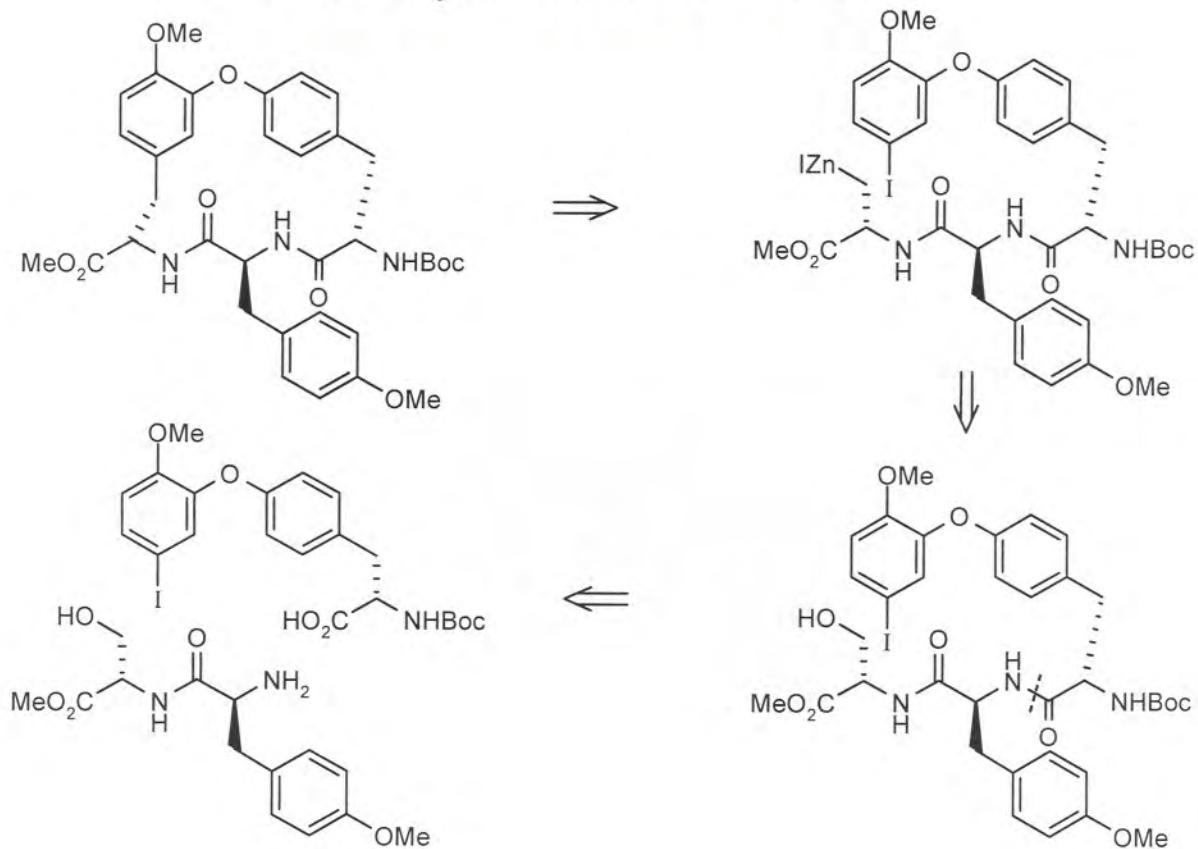
Comparison of ^{13}C Data with Literature data

Pearson

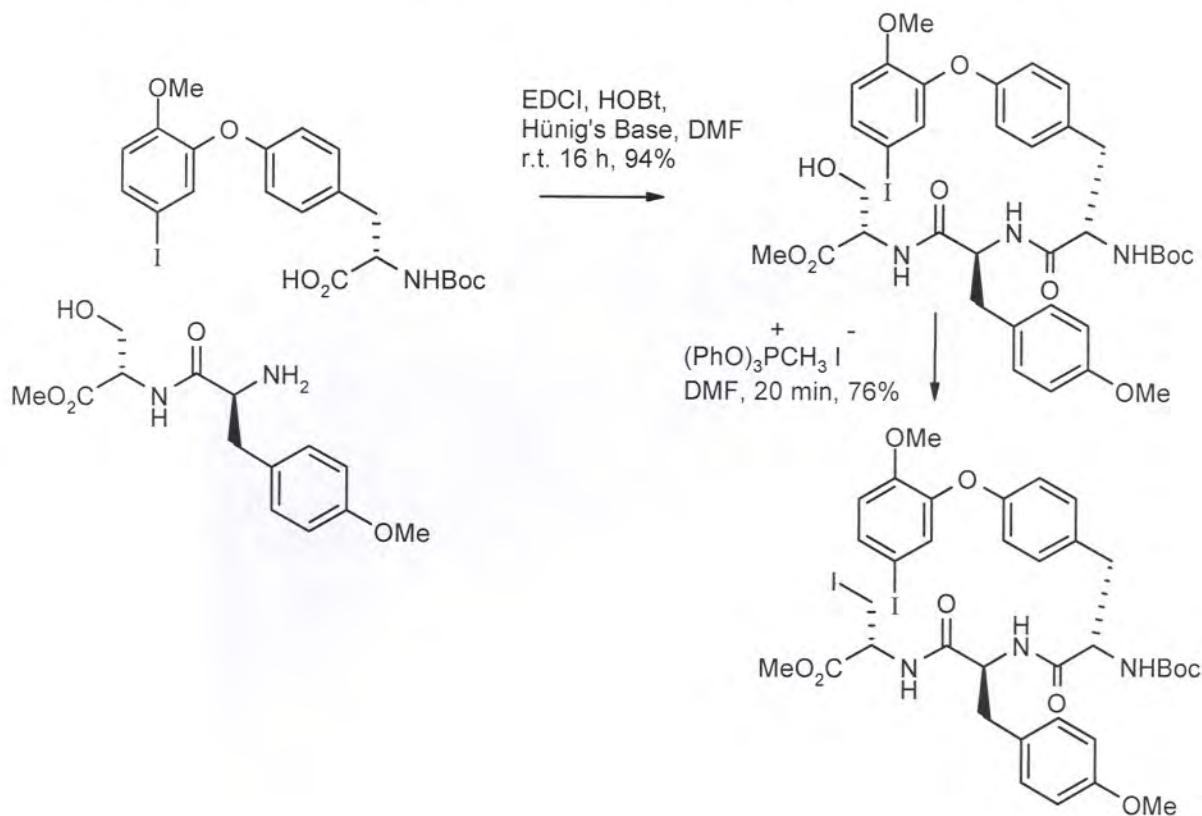




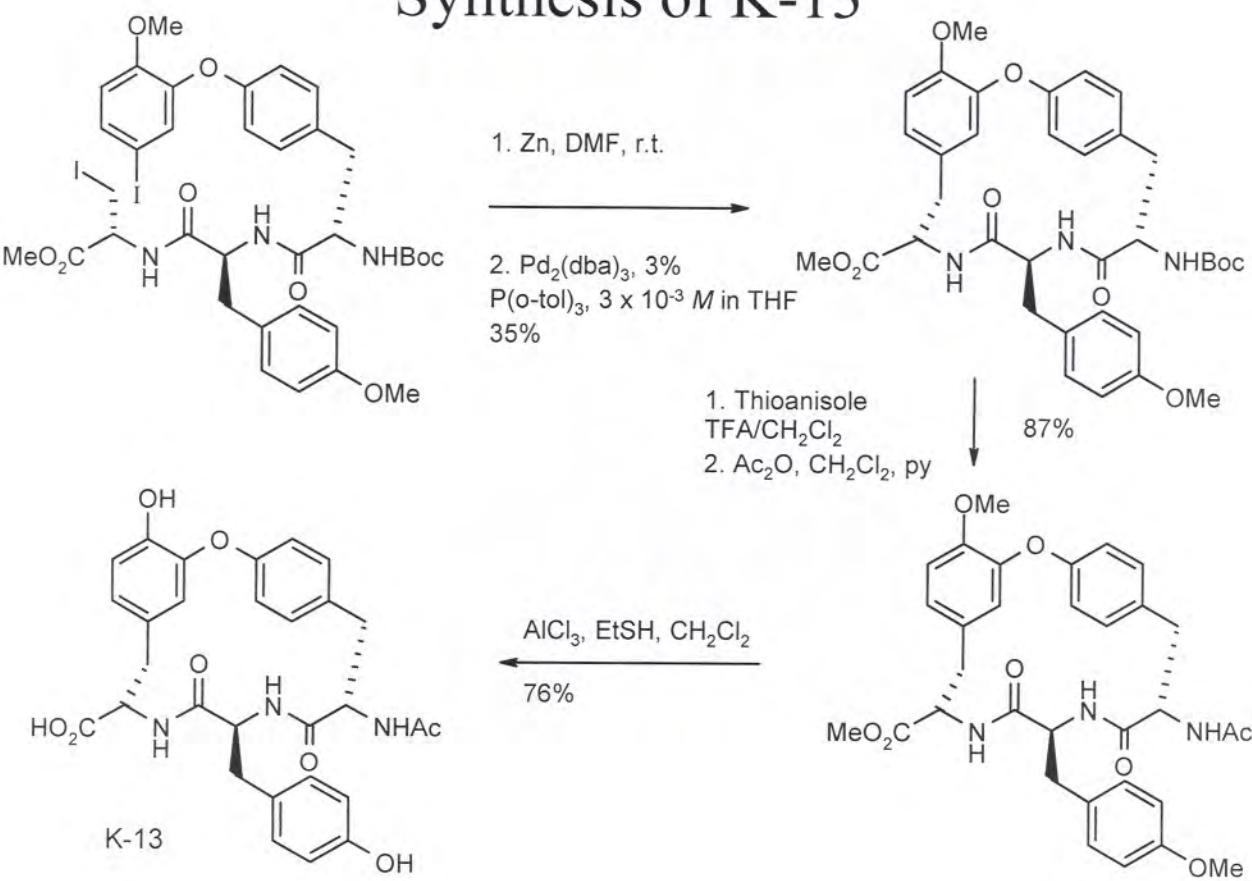
Retrosynthesis for K-13



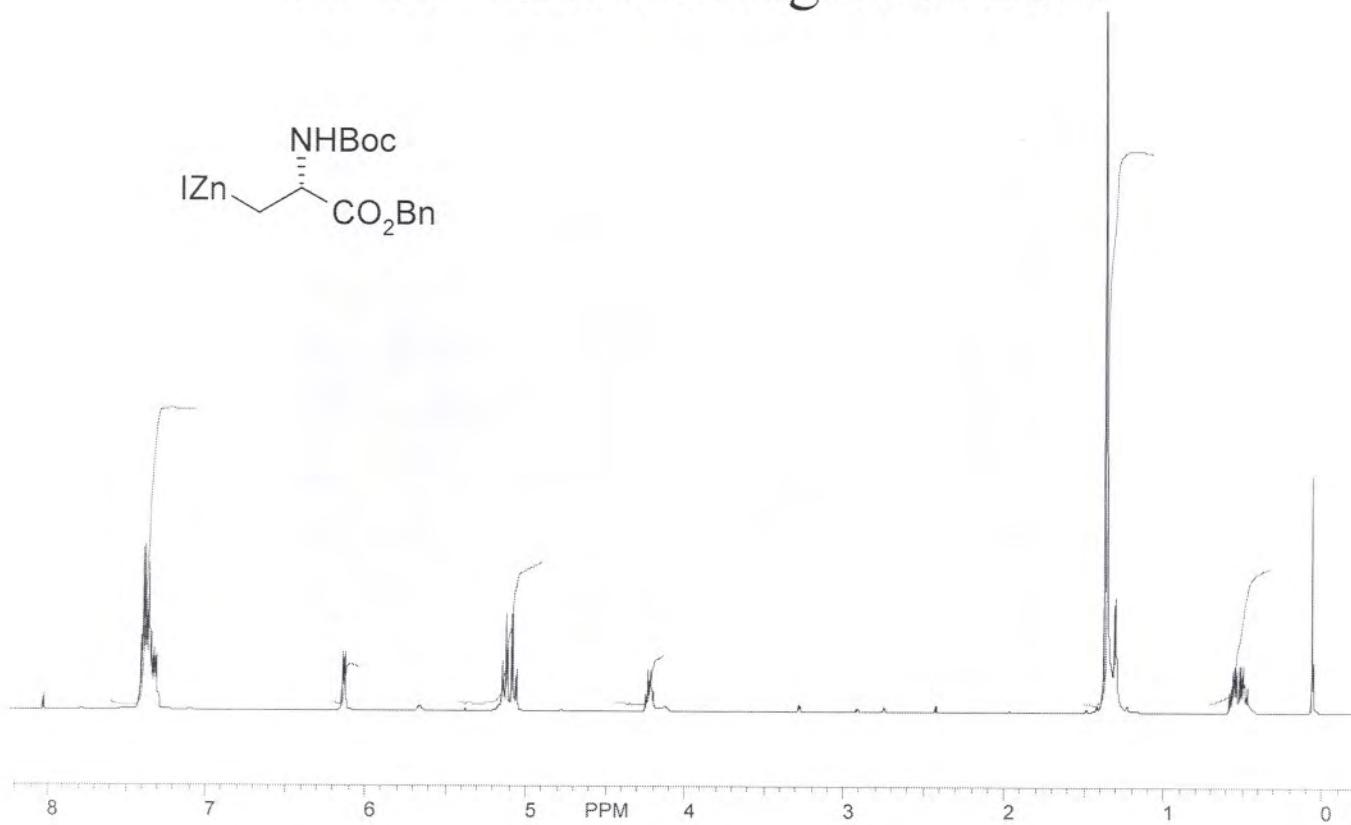
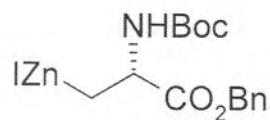
Synthesis of the Macrocyclic Precursor



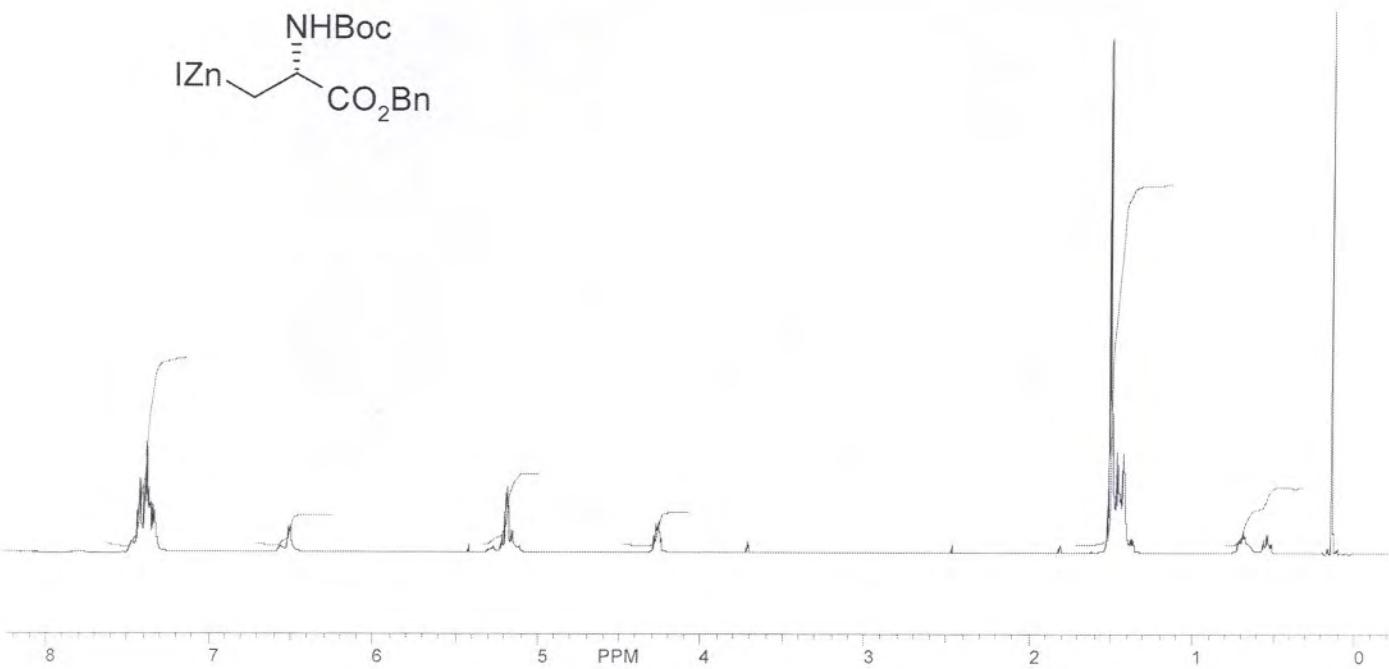
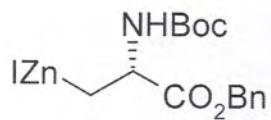
Synthesis of K-13



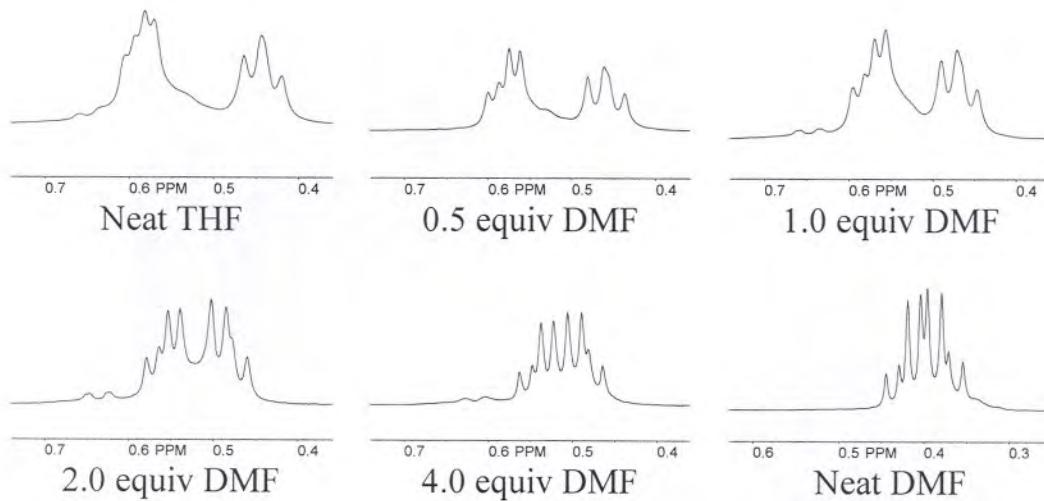
Serine-Derived Reagent in DMF



Serine-Derived Reagent in THF



Solvent Dependence of NMR Spectra



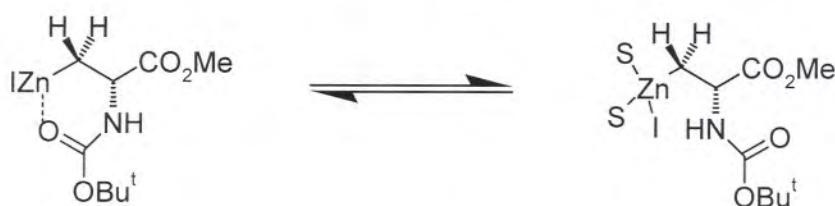
Interpretation of ^1H NMR data

THF

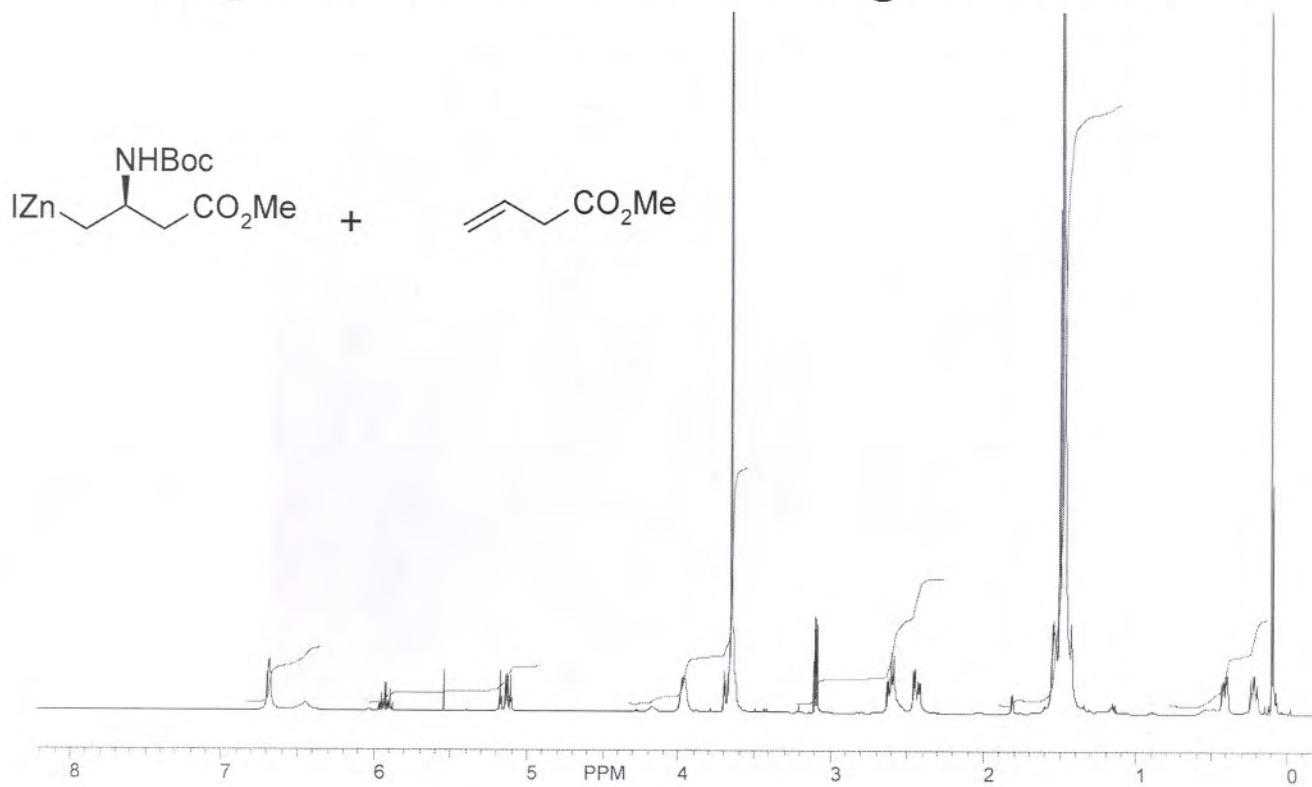
- Broad signals
- Large chemical shift difference
- Different values for J_{AX} and J_{BX}

DMF

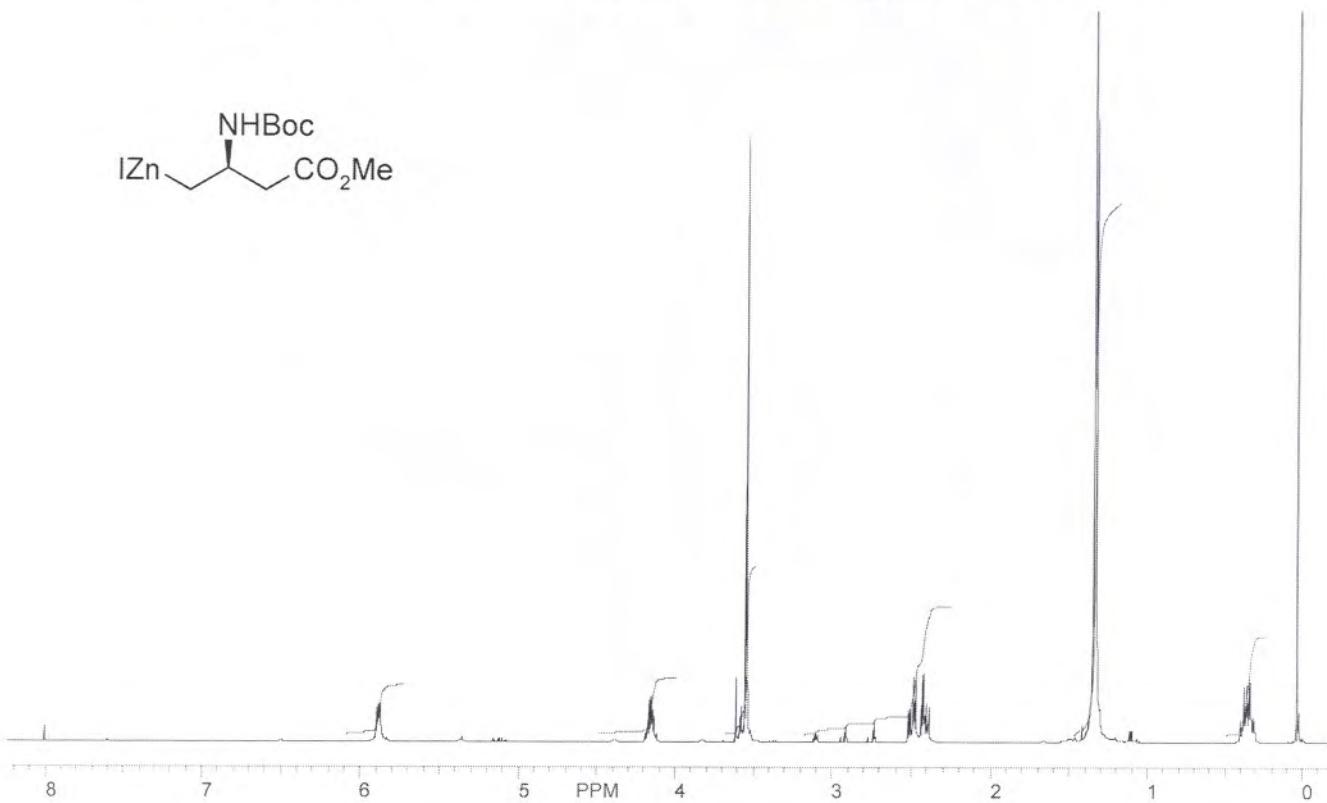
- Sharp signals
- Small chemical shift difference
- Similar values for J_{AX} and J_{BX}



Aspartic Acid-Derived Reagent in THF



Aspartic Acid-Derived Reagent in DMF

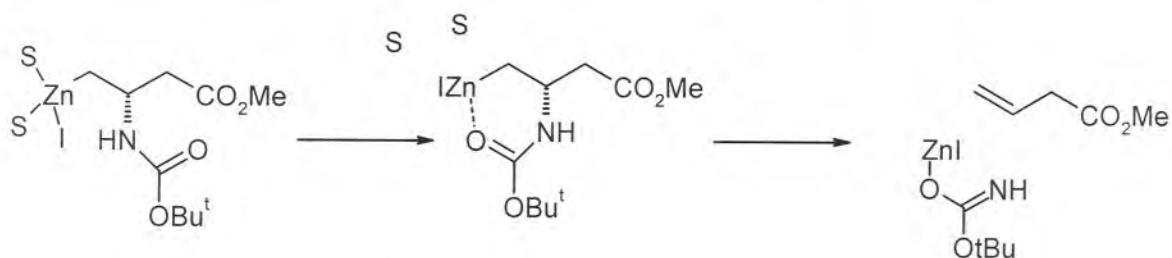


Kinetics of Decomposition



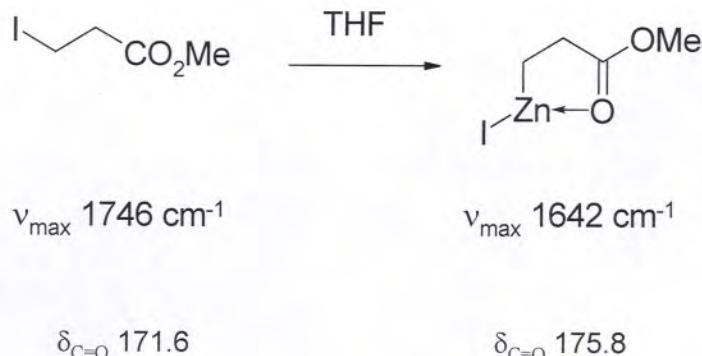
- Faster decomposition in THF than in DMF ΔH^\ddagger ΔS^\ddagger
kJ mol⁻¹ J K⁻¹ mol⁻¹
- Simple first order kinetics in both THF and DMF THF-d₈ + 70 - 85
- Large negative ΔS^\ddagger in THF indicating a highly ordered transition state DMF-d₇ + 90 - 31

Tentative Mechanism for Elimination



- Elimination proceeds *via* coordination of carbamate group to zinc, to give a highly ordered transition structure, hence the negative *entropy* of activation.
- THF interacts more weakly with the zinc than does DMF, as indicated by ¹H NMR. Thus, the disruption of DMF coordination to zinc in forming the transition structure also accounts for the higher *enthalpy* of activation in this solvent.
- The release of coordinated DMF partially compensates for the formation of the highly ordered transition structure, so the *entropy* of activation is less negative in this solvent.

Internal Coordination in Functionalised Organozinc Halides



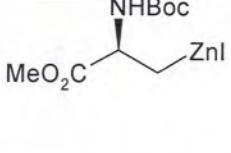
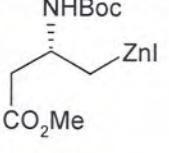
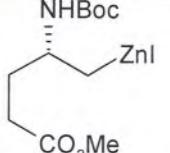
^{13}C NMR Data

- Chemical shift differences of the carbonyl groups of the β -amino zinc reagents in d_8 -THF relative to the parent iodides.

$\Delta\delta (\delta_{(\text{R-ZnI})} - \delta_{(\text{R-I})})$			
Ester	+5.347	+1.464	+ 0.675
Carbamate	+ 2.711	+3.747	+3.923

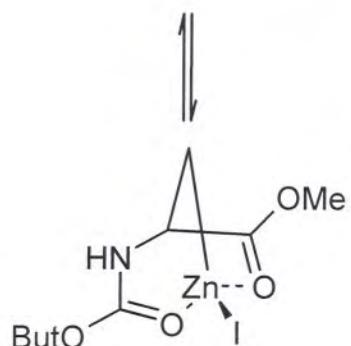
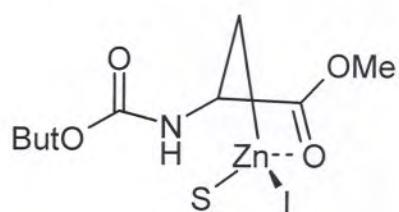
^{13}C NMR Data

- Chemical shift differences of the carbonyl groups of the β -amino zinc reagents in $d_7\text{-DMF}$ relative to the parent iodides.

$\Delta\delta$ ($\delta_{(\text{R-ZnI})} - \delta_{(\text{R-I})}$)			
			
Ester	+5.786	+0.872	+0.115
Carbamate	-0.868	-0.535	-0.687

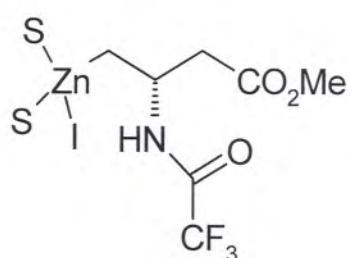
Unique Stability of Serine-Derived Reagent

- For the serine-derived reagent in all solvents, co-ordination of ester to zinc occurs.

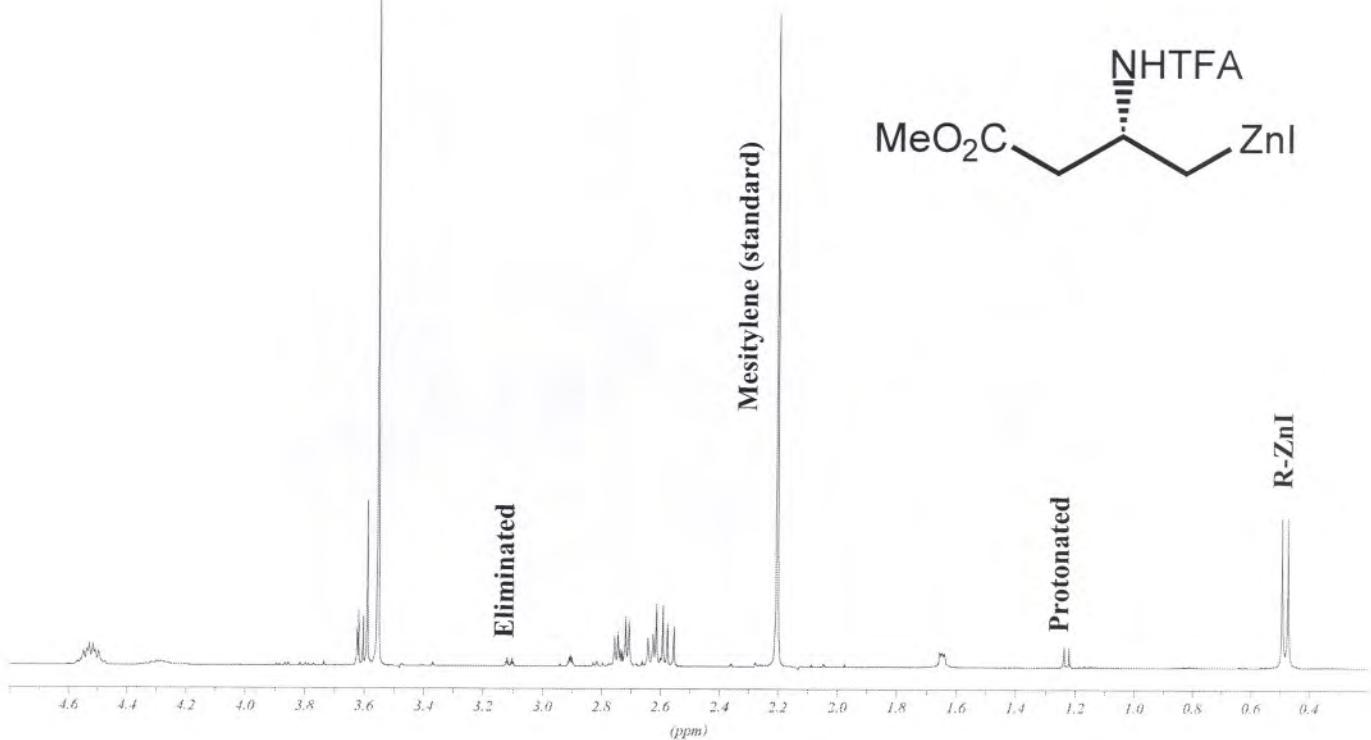


How to Minimise the Elimination

- The elimination reaction appears to be dependent on the Lewis basicity of the carbonyl function, not its ability as a leaving group.
- Therefore, we need to choose a group in which Lewis basicity is minimised.

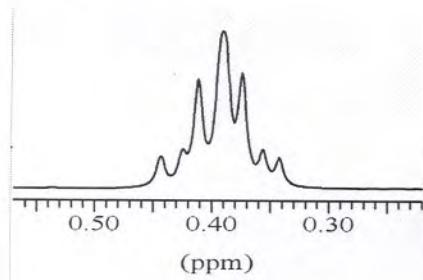


N-TFA Asp(OMe)-ZnI in d₇-DMF

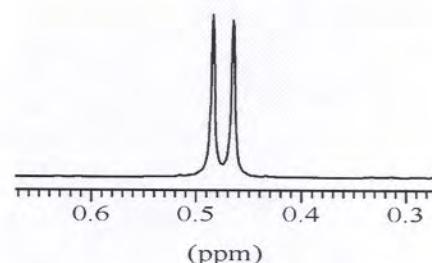


Comparison of CH_2ZnI signals

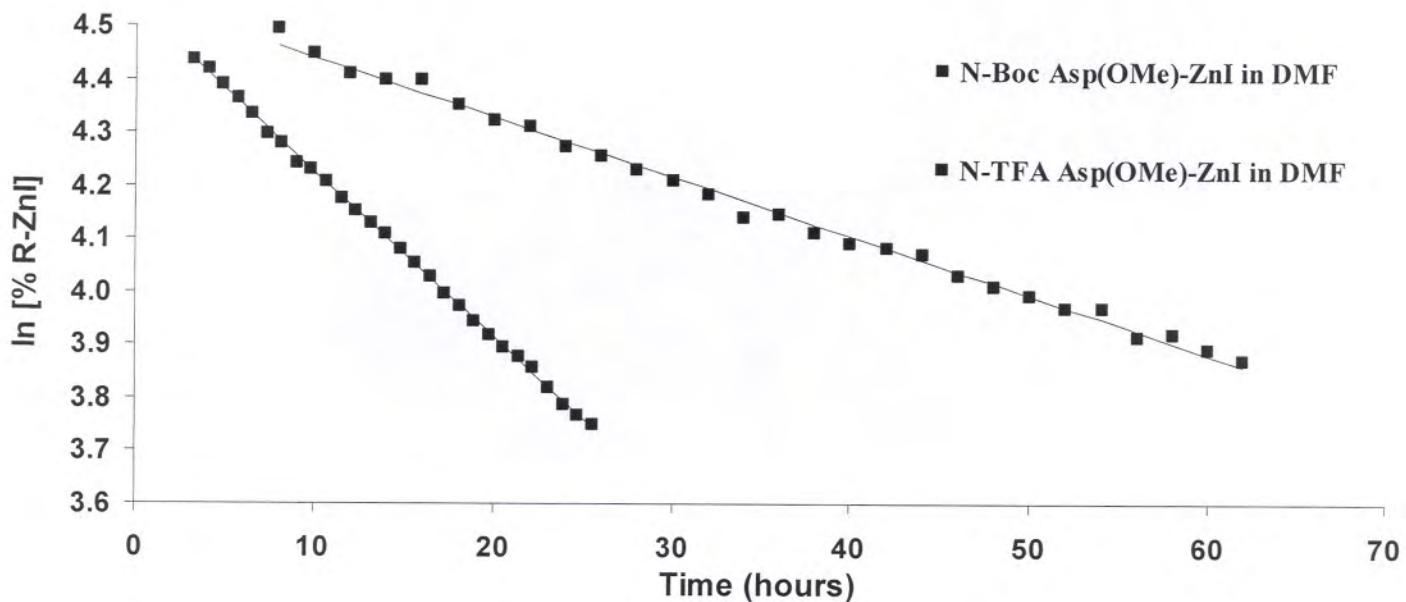
Methylene protons for
N-Boc Asp(OMe)-ZnI
in DMF



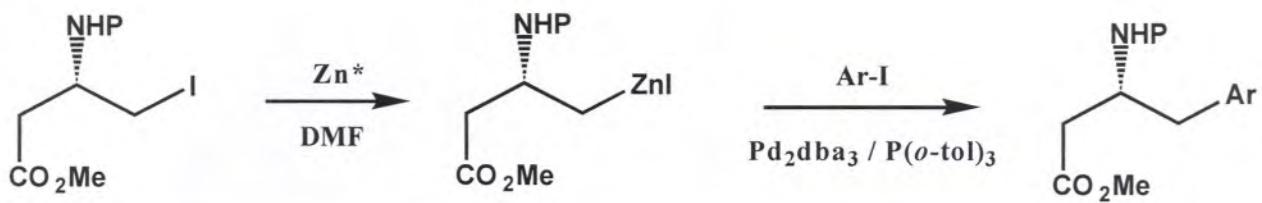
Methylene protons for
N-TFA Asp(OMe)-ZnI
in DMF



Decomposition of Organozinc Reagents



Yields from Palladium-Catalysed Cross Coupling of β -Amidozinc Reagents



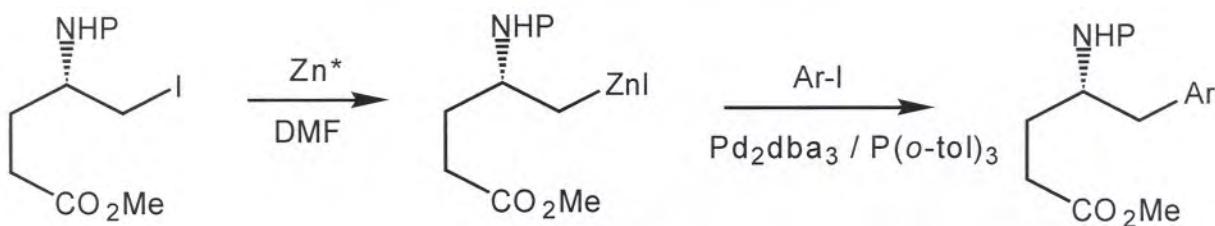
Ar	P = Boc	P = TFA
	Yield (%)	Yield (%)
4-Me-Ph	73	64
4-MeO-Ph	68	69
Ph	73	72
4-CN-Ph	-	77
1-Naphthyl	61	70
4-Br-Ph	58	53

Jackson, R. F. W.; Rilatt, I. and Murray, P. J., *Chem. Commun.*, 2003, 1242 – 1243.

Kinetic Behaviour of Organozinc Reagents

Organozinc Reagent	Rate of Decomposition ($\times 10^{-5}$ s $^{-1}$)	Rate of Coupling ($\times 10^{-4}$ s $^{-1}$)	Isolated Yield (%)
	0.87	Too fast to measure	63
	0.31	20.6	65
	2.84	41.3	70
	0.33	62.5	88

Palladium-Catalysed Cross Coupling of β -Amidozinc Reagents

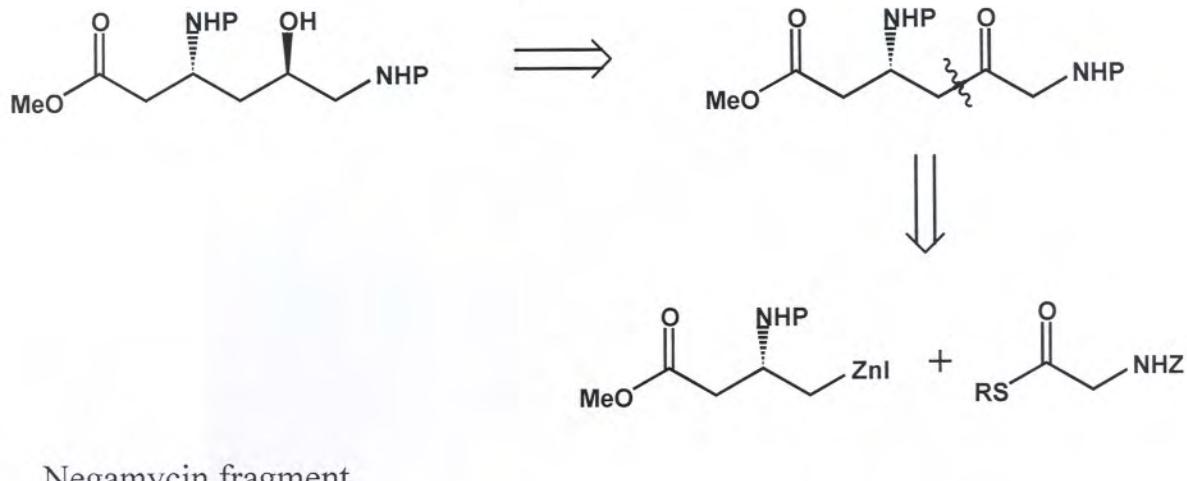


Ar	Yield with Boc (%)	Yield with TFA (%)
4-MeO-Ph	68	79
2-NH ₂ -Ph	56	62
2-F-Ph	34	51
4-Me-Ph	68	79
Ph	68	88

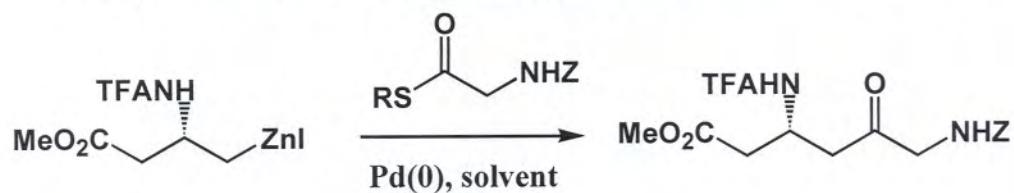
Reactivity of Organozinc Reagents

Organozinc Reagent	Rate of Decomposition ($\times 10^{-5}$ s ⁻¹)	Rate of Coupling ($\times 10^{-5}$ s ⁻¹)	Isolated Yield (%)
	1.1	Too fast to measure	74
	0.42	Too fast to measure	83
	0.2	8.5	44

Applications of Reagents



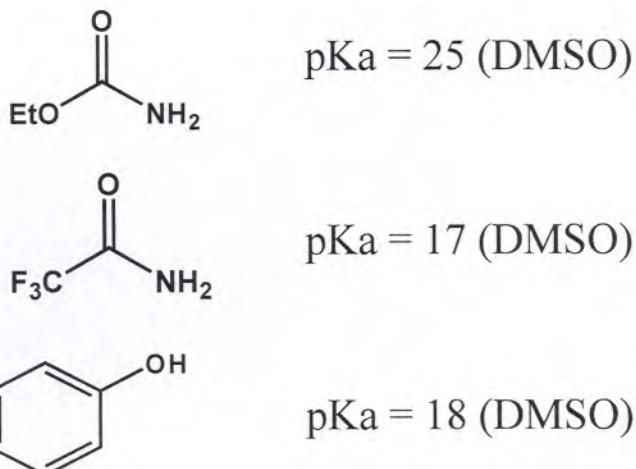
Kinetics of Coupling Reaction – Thioesters



R	Solvent	Reaction Rate ($\times 10^{-5} \text{ s}^{-1}$)	Reaction Time	Isolated Yield (%)
Ph	DMF	0.5	> 16 h	65
Ph	THF/DMF(6 eq)	1.0	~ 16 h	72
C ₆ F ₅	DMF	1.7	~ 8 h	81
C ₆ F ₅	THF/DMF(6 eq)	5.2	~ 3 h	94

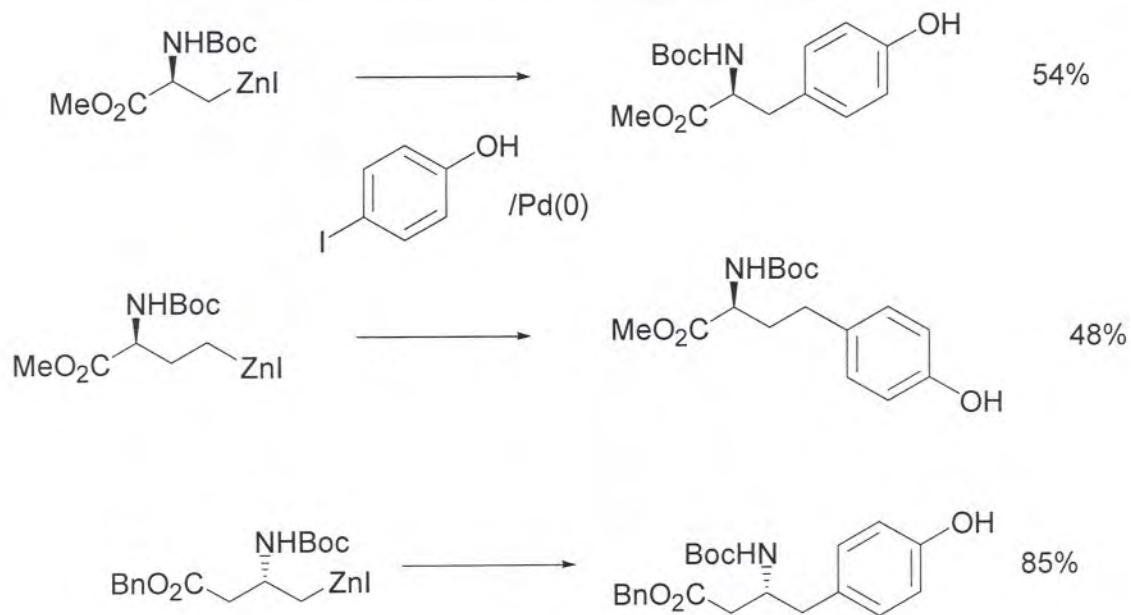
Conclusions

- TFA reagents more stable to elimination
- Ester coordination critical factor in reactivity
- Glutamic acid trifluoroacetamide a superior reagent
- More acidic proton of the trifluoroacetamide tolerated



F. G. Bordwell, *Acc. Chem. Res.*, 1988, **21**, 456-463

Coupling with Iodophenols



Org. Biomol. Chem., 2004, **2**, 110 – 113.

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